

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Physical Science Chapter 6 Chemical Bonds Test Review

D Keegan



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Table of Contents Physical Science Chapter 6 Chemical Bonds Test Review

1. Understanding the eBook Physical Science Chapter 6 Chemical Bonds Test Review
 - The Rise of Digital Reading Physical Science Chapter 6 Chemical Bonds Test Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 6 Chemical Bonds Test Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 6 Chemical Bonds Test Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 6 Chemical Bonds Test Review
 - Personalized Recommendations
 - Physical Science Chapter 6 Chemical Bonds Test Review User Reviews and Ratings
 - Physical Science Chapter 6 Chemical Bonds Test Review and Bestseller Lists
5. Accessing Physical Science Chapter 6 Chemical Bonds Test Review Free and Paid eBooks
 - Physical Science Chapter 6 Chemical Bonds Test Review Public Domain eBooks
 - Physical Science Chapter 6 Chemical Bonds Test Review eBook Subscription Services
 - Physical Science Chapter 6 Chemical Bonds Test Review Budget-Friendly Options

6. Navigating Physical Science Chapter 6 Chemical Bonds Test Review eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 6 Chemical Bonds Test Review Compatibility with Devices
 - Physical Science Chapter 6 Chemical Bonds Test Review Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 6 Chemical Bonds Test Review
 - Highlighting and Note-Taking Physical Science Chapter 6 Chemical Bonds Test Review
 - Interactive Elements Physical Science Chapter 6 Chemical Bonds Test Review
8. Staying Engaged with Physical Science Chapter 6 Chemical Bonds Test Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 6 Chemical Bonds Test Review
9. Balancing eBooks and Physical Books Physical Science Chapter 6 Chemical Bonds Test Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 6 Chemical Bonds Test Review
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter 6 Chemical Bonds Test Review
 - Setting Reading Goals Physical Science Chapter 6 Chemical Bonds Test Review
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter 6 Chemical Bonds Test Review
 - Fact-Checking eBook Content of Physical Science Chapter 6 Chemical Bonds Test Review
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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