

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

Na, K, S, O, Cl

3. Which atom has a higher first ionization energy than sodium?

Chlorine (because it has a higher electronegativity and is closer to fluorine)

4. Which two elements in the same family generally have similar properties?

Sulfur and chlorine (because they are both in the same period and have similar properties)

5. Arrange the following elements in order of increasing atomic radius from left to right: sodium, potassium, calcium.

- a. sodium (because it has the lowest electronegativity) **Na**
 b. potassium (because it has the lowest electronegativity) **K**
 c. calcium (because it has the highest electronegativity) **Ca**

6. Which period in the periodic table shows atoms with the highest electronegativity? Which element has the highest?

Period 2, Fluorine (because it has the highest electronegativity)

7. Which period in the periodic table shows atoms with the lowest electronegativity? Which element has the lowest?

Period 7, Francium (because it has the lowest electronegativity)

8. Write the atomic radius of the following elements in order of increasing atomic radius.

Fr, Cs, Rb, K, Na, Li

Fr, Cs, Rb, K, Na, Li

9. Write the atomic radius of the following elements in order of increasing atomic radius.

Fr, Cs, Rb, K, Na, Li

Fr, Cs, Rb, K, Na, Li

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