

Group 15 Elements

N, P, As, Sb, Bi, Mc

Hydrides of Group 15 Elements

Halides of Group 15 Elements

The valence shell electronic configuration is ns^2np^3

N and P - non-metals

As and Sb - metalloids

Bi and Mc (moscovium) - metals

Common oxidation states are -3, +3 and +5

NH_3 has higher melting and boiling points than PH_3 due to hydrogen bonding

NH_3 is a mild reducing agent and BiH_3 is the strongest reducing agent

Basicity order - $NH_3 > PH_3 > AsH_3 > SbH_3 > BiH_3$

Pentahalides are more covalent than trihalides

Nitrogen does not form pentahalide due to absence of the d-orbitals in its valence shell

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