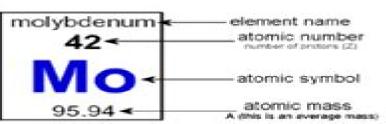
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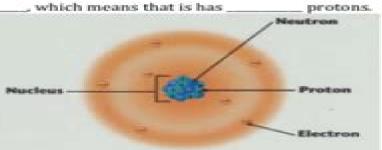
SPS1: Students will investigate our current understanding of the atom.

a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
 - Proton has a charge of ______ and is located in the ______ of the atom.
 - Protons are also the elements ______ number of the identity of the element, so protons are found by looking at the atomic number
 - Neutron has a charge of ______ and is located in the ______ of the atom.
 - Neutrons are found by subtracting the _____ Number from the _____ Mass or the BIG #- SMALL #.
 - Electron has a charge of ______ and is located in the ______ of the atom.
 - Electrons are equal to the number of ______ in a neutral atom.
 - The nucleus is made up of these 2 particles _____ and _____
 - The electrons are located in this region of the atom, known as the ______
 - The electron cloud is divided in to energy levels or energy shells, the first level will hold electrons and the second level will hold ______ electrons.
 - The valence shell or valence level is the ______ energy shell of the atom.
- atomic mass and atomic number.
 - Atomic number is equal to the number of ______ in the atom and is also the atoms identity—it's the small number of the periodic table.
 - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
 - Define Isotope__
 - Ex. Carbon-12 and Carbon-14
 - Carbon 12 has ____P ___N ___E
 - Carbon 14 has ____P ___N ___E
- explain the relationship of the proton number to the element's identity.
 - The _____ number is equal to the number of _____ in the atom
 which is how the element is identified.
 - Ex. Nitrogen has an atom number of ______, which means that is has ______ protons.



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number Atomic Number



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