

# Common Polyatomic Ions

|                      |                                       |
|----------------------|---------------------------------------|
| acetate              | $\text{C}_2\text{H}_3\text{O}_2^-$    |
| ammonium             | $\text{NH}_4^+$                       |
| arsenate             | $\text{AsO}_4^{3-}$                   |
| arsenite             | $\text{AsO}_3^{3-}$                   |
| azide                | $\text{N}_3^-$                        |
| benzoate             | $\text{C}_7\text{H}_5\text{O}_2^-$    |
| borate               | $\text{BO}_3^{3-}$                    |
| bromate              | $\text{BrO}_3^-$                      |
| carbonate            | $\text{CO}_3^{2-}$                    |
| chlorate             | $\text{ClO}_3^-$                      |
| chlorite             | $\text{ClO}_2^-$                      |
| chromate             | $\text{CrO}_4^{2-}$                   |
| cyanide              | $\text{CN}^-$                         |
| dichromate           | $\text{Cr}_2\text{O}_7^{2-}$          |
| dihydrogen phosphate | $\text{H}_2\text{PO}_4^-$             |
| dihydrogen phosphite | $\text{H}_2\text{PO}_3^-$             |
| hydrogen carbonate   | $\text{HCO}_3^-$                      |
| hydrogen phosphate   | $\text{HPO}_4^{2-}$                   |
| hydrogen phosphite   | $\text{HPO}_3^{2-}$                   |
| hydrogen sulfate     | $\text{HSO}_4^-$                      |
| hydrogen sulfide     | $\text{HS}^-$                         |
| hydrogen sulfite     | $\text{HSO}_3^-$                      |
| hydroxide            | $\text{OH}^-$                         |
| hypochlorite         | $\text{ClO}^-$                        |
| iodate               | $\text{IO}_3^-$                       |
| manganate            | $\text{MnO}_4^{2-}$                   |
| nitrate              | $\text{NO}_3^-$                       |
| nitrite              | $\text{NO}_2^-$                       |
| oxalate              | $\text{C}_2\text{O}_4^{2-}$           |
| perchlorate          | $\text{ClO}_4^-$                      |
| permanganate         | $\text{MnO}_4^-$                      |
| peroxide             | $\text{O}_2^{2-}$                     |
| phosphate            | $\text{PO}_4^{3-}$                    |
| phosphite            | $\text{PO}_3^{3-}$                    |
| silicate             | $\text{SiO}_3^{2-}$                   |
| sulfate              | $\text{SO}_4^{2-}$                    |
| sulfite              | $\text{SO}_3^{2-}$                    |
| tartrate             | $\text{C}_4\text{H}_4\text{O}_6^{2-}$ |
| thiocyanate          | $\text{SCN}^-$                        |
| thiosulfate          | $\text{S}_2\text{O}_3^{2-}$           |

|                                       |                      |
|---------------------------------------|----------------------|
| $\text{AsO}_3^{3-}$                   | arsenite             |
| $\text{AsO}_4^{3-}$                   | arsenate             |
| $\text{BO}_3^{3-}$                    | borate               |
| $\text{BrO}_3^-$                      | bromate              |
| $\text{C}_2\text{H}_3\text{O}_2^-$    | acetate              |
| $\text{C}_2\text{O}_4^{2-}$           | oxalate              |
| $\text{C}_4\text{H}_4\text{O}_6^{2-}$ | tartrate             |
| $\text{C}_7\text{H}_5\text{O}_2^-$    | benzoate             |
| $\text{ClO}^-$                        | hypochlorite         |
| $\text{ClO}_2^-$                      | chlorite             |
| $\text{ClO}_3^-$                      | chlorate             |
| $\text{ClO}_4^-$                      | perchlorate          |
| $\text{CN}^-$                         | cyanide              |
| $\text{CO}_3^{2-}$                    | carbonate            |
| $\text{Cr}_2\text{O}_7^{2-}$          | dichromate           |
| $\text{CrO}_4^{2-}$                   | chromate             |
| $\text{H}_2\text{PO}_3^-$             | dihydrogen phosphite |
| $\text{H}_2\text{PO}_4^-$             | dihydrogen phosphate |
| $\text{HCO}_3^-$                      | hydrogen carbonate   |
| $\text{HPO}_3^{2-}$                   | hydrogen phosphite   |
| $\text{HPO}_4^{2-}$                   | hydrogen phosphate   |
| $\text{HS}^-$                         | hydrogen sulfide     |
| $\text{HSO}_3^-$                      | hydrogen sulfite     |
| $\text{HSO}_4^-$                      | hydrogen sulfate     |
| $\text{IO}_3^-$                       | iodate               |
| $\text{MnO}_4^-$                      | permanganate         |
| $\text{MnO}_4^{2-}$                   | manganate            |
| $\text{N}_3^-$                        | azide                |
| $\text{NH}_4^+$                       | ammonium             |
| $\text{NO}_2^-$                       | nitrite              |
| $\text{NO}_3^-$                       | nitrate              |
| $\text{O}_2^{2-}$                     | peroxide             |
| $\text{OH}^-$                         | hydroxide            |
| $\text{PO}_3^{3-}$                    | phosphite            |
| $\text{PO}_4^{3-}$                    | phosphate            |
| $\text{S}_2\text{O}_3^{2-}$           | thiosulfate          |
| $\text{SCN}^-$                        | thiocyanate          |
| $\text{SiO}_3^{2-}$                   | silicate             |
| $\text{SO}_3^{2-}$                    | sulfite              |
| $\text{SO}_4^{2-}$                    | sulfate              |

# Physical Science Formulas With Polyatomic Ions Answer

**F Rizvi**



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