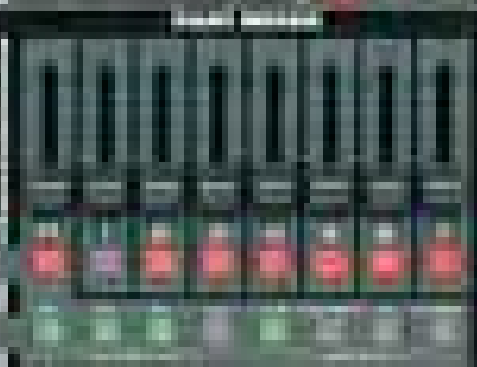
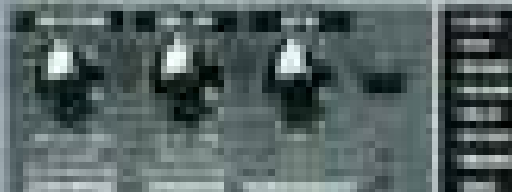
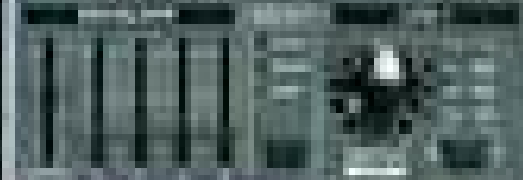
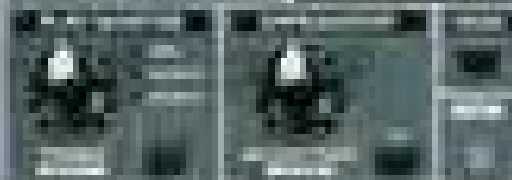
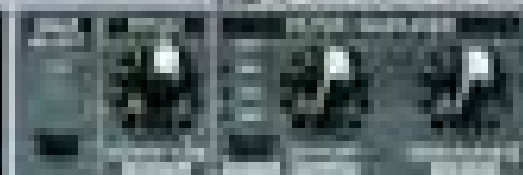


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web solution $v_1 = 250\text{ l}$ $v_2 = 273\text{ K}$ 303 K $t_2 = 273\text{ K}$ 343 K we will be applying charles law as pressure is constant here $v_1/t_1 = v_2/t_2$ or $v_2 = v_1 \cdot t_2/t_1 = 250\text{ l} \cdot 303\text{ K}/273\text{ K} = 283\text{ l}$ at 90°C a helium sample has a volume of 500 ml determine the temperature at which the volume of the liquid will become 240 ml