

Relative Formula Mass

If H_2O is the chemical formula for water, how many atoms are in total in the substance? **3**

In chemical symbols each capital letter indicates the next element. Use a periodic table to complete the following. For the following substance list the names of the elements and the number of atoms present in the chemical formula.

KNO_3 _____

Na_2CO_3 _____

The relative atomic mass, A_r , is the mean mass of an atom of an element compared with one twelfth of

the mass of a _____ 12 atom.

Circle on the Calcium tile the relative atomic mass.



Relative formula mass is the mean mass of a substance if all the relative atomic masses are _____ together. The relative atomic mass of carbon is 12 and oxygen is 16. The relative formula mass of CO_2 is $12 + (2 \times 16) = 44$. Complete the following (Use a periodic table):

CH_4 _____

$MgCl_2$ _____

$Mg(NO_3)_2$ _____

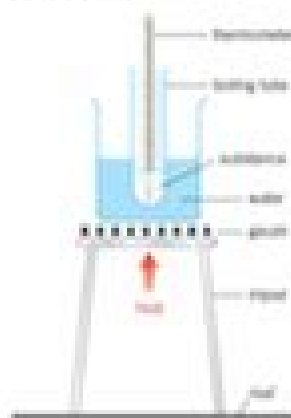
C2.1 - Separating Mixtures - Combined Chemistry Revision Sheet

Pure and Impure Substances

Complete the sentences using the words in the box. Use each word once, more than once or not at all.

metals steel mineral mixed harder impure one melting many iron pure joined distilled alloys sulfur

In science a pure substance consists of _____ element/s or compound/s. Mixtures are _____. They contain two or more different elements or compounds _____ together _____. water and _____ are mixtures. Mixtures of metals, called _____ are _____ than pure metals. We often determine purity of a substance by looking at its _____ point.



Three students, A, B and C make aspirin and check its purity. The students test the purity of the aspirin by measuring the melting point, using the equipment shown. Their results are shown in the table.

What two things do they need to record to draw a change of state graph?

What do they need to do during the experiment to improve accuracy?

Student	Melting point ($^{\circ}C$)
A	140
B	139
C	137

The melting point of pure aspirin is in the range $138-140^{\circ}C$. Which student made pure aspirin?

Filtering

Label the diagram below using the labels in the box once.

funnel filtrate beaker
conical flask filter paper
solution residue



Answer the following questions about when substances dissolve.

When a substance dissolves in another substance a _____ is made. A substance which can be dissolved is said to be _____.

The substance that dissolves is called the _____.

The substance it dissolves in is called the _____.

When a substance does not dissolve it is said to be _____.

Empirical Formula

Empirical formula shows the simplest whole number ratio of atoms of each element in a compound.

Look at the space saving models of Butane.

Number of carbon atoms _____

Number of hydrogen atoms _____

Chemical formula _____ Empirical formula _____

Look at the displayed formula of valeric acid.

Number of carbon atoms _____

Number of hydrogen atoms _____

Number of oxygen atoms _____

Chemical formula _____ Empirical formula _____

Derive the empirical formula for the following:

Al_2O_3 _____ $C_4H_8O_2$ _____

$Ca_3(PO_4)_2$ _____ Na_2SO_4 _____

Crystallisation

Number the following steps to show the order you would carry them out to produce a sample of pure copper sulfate crystals. The first step has been done for you.



Label the diagram below using the labels in the box once.

Bunsen burner tripod
gauze evaporating basin

- ☐ Add an excess of copper carbonate powder to the acid and stir gently with a stirring rod.
- ☐ Carefully warm the solution of copper sulfate so that the water evaporates.
- ☐ Stop heating when there is a teaspoon of liquid left, allow the copper sulfate to cool and crystallise fully.
- ☒ 1 Place 25 cm³ sulfuric acid solution into a beaker.
- ☐
- ☐ Fold a filter paper and place it into the filter funnel.

Carefully pour the mixture through the filter paper to remove excess copper carbonate powder. The filtrate passes through the filter paper and is

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Sebastian Brünink

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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