

phosphorus(V) oxide with water

phosphoric (V) acid

$P_4O_{10}$  → single molecular  
→ van der Waals

$MgO$  → giant ionic  
→ ionic

[2]

- (d) Explain why phosphorus(V) oxide has a low melting point of approximately  $300^\circ\text{C}$  but magnesium oxide has a high melting point of approximately  $2850^\circ\text{C}$ .

The melting of phosphorus(V) oxide involve the breaking of weak van der Waals' forces between molecules.

The melting of magnesium oxide involve the breaking of electrostatic forces between cation and anion which is ionic bonds.

The ionic bonds is stronger than vdw forces.

[3]

solved by TheamKL

# October November 2chemistry Paper 606

**B Lingard**



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