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g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield
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c is the limiting reagent 4 33 g of h 2 are left over 13

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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of co is
reacted with 8.6 g of H_2

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problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical
reaction

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web limiting reactant and percent yield practice name 1 consider the following reaction $\text{nh}_4\text{no}_3 + \text{na}_3\text{po}_4 \rightarrow \text{nh}_4\text{}_3\text{po}_4 + \text{nano}_3$ which reactant is limiting assuming we started with 30 0 grams of ammonium nitrate and 50 0 grams of sodium phosphate what is the

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3