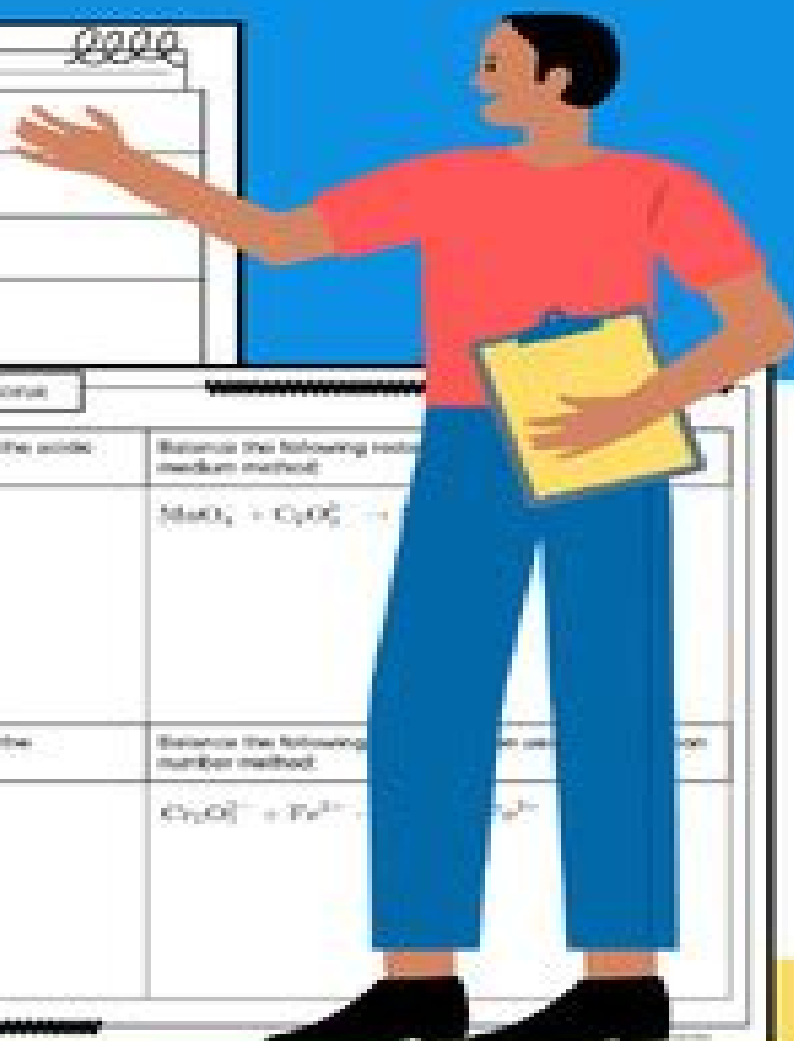


OXIDATION-REDUCTION REACTIONS

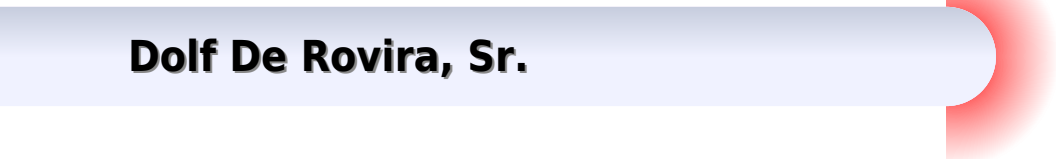
Oxidation Reactions in Everyday Life			
		Name _____	
Burning of iron is an example of a redox reaction.	TRUE	FALSE	Justify:
Combustion reactions only occur in the presence of oxygen.	TRUE	FALSE	Justify:
Corrosion is the process of metal gaining electrons.	TRUE	FALSE	Justify:
Rusting of iron can be accelerated by exposure to sulfuric acid.	TRUE	FALSE	Justify:
Combustion reactions release energy in the form of heat and light.	TRUE		
Corrosion is a slow process that occurs in metals.	TRUE		
Combustion reactions involve the reduction of oxygen.	TRUE		
The rusting of iron is an example of a non-redox reaction.	TRUE		

Steps for Balancing Redox Equations			
Balance the following redox equation using the oxidation number method.		Balance the following redox equation using the oxidation number method.	
$\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$		$\text{MnO}_2 + \text{C}_2\text{O}_4^{2-} \rightarrow$	
Balance the following redox equation using the half-reaction method.		Balance the following redox equation using the half-reaction method.	
$\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$		$\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$	



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