

Chemistry Chapter 8 Test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ____ 1. What type of bond would you expect in a compound of oxygen and sodium?
☒ a. ionic
 b. metallic
 c. covalent
 d. polar covalent
- ____ 2. The symbols δ^- and δ^+ indicate which of the following?
 a. electronegativity
 b. electrostatic force
 c. charge
☒ d. partial charge
- ____ 3. Which of the following represents a nonpolar covalent molecule?
 a. NaCl
☒ b. Br₂
 c. HCN
 d. Au
- ____ 4. The measure of an atom's ability to attract electrons to itself in a chemical bond is called _____.
 a. electron affinity
 b. polarity
☒ c. electronegativity
 d. ionization energy
- ____ 5. The polarity of a bond depends on the difference in the atoms' _____.
 a. electron affinities
☒ b. electronegativities
 c. atomic radii
 d. ionization energies
- ____ 6. What type of bond would you expect to form between nitrogen and potassium?
☒ a. ionic
 b. covalent
 c. polar covalent
 d. metallic
- ____  7. High electronegativity differences are found in _____ bonds.
 a. metallic
☒ b. ionic metal - nonmetal
 c. covalent
 d. Pauling
- ____  8. To achieve an octet when combining with ¹potassium, would an atom of a Group ⁶16 element be expected to lose, gain, or share electrons?
 a. lose
☒ b. gain
 c. share
☒ d. This cannot be determined from the given information. x
- ____ 9. Which of the following does not exist as a diatomic molecule?
☒ a. argon
 b. hydrogen
 c. chlorine
 d. oxygen
 H N O F Cl Br I
 Gas : H₂ N₂ O₂ F₂ Cl₂ Br₂ I₂
- ____ 10. Which of these is not a covalent molecule?
 a. F₂ =
 b. CO =
☒ c. LiF
 d. NH₃
- ____ 11. A dash in a Lewis structure represents _____.
☒ a. one electron
 b. two electrons
 c. four electrons
 d. unpaired electrons
- ____ 12. The central atom in a Lewis structure is usually the atom _____.

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1. Understanding the eBook Odysseyware Chemistry Test Answers
 - The Rise of Digital Reading Odysseyware Chemistry Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Odysseyware Chemistry Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Odysseyware Chemistry Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Odysseyware Chemistry Test Answers
 - Personalized Recommendations

- Odysseyware Chemistry Test Answers User Reviews and Ratings
- Odysseyware Chemistry Test Answers and Bestseller Lists
- 5. Accessing Odysseyware Chemistry Test Answers Free and Paid eBooks
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 - Odysseyware Chemistry Test Answers eBook Subscription Services
 - Odysseyware Chemistry Test Answers Budget-Friendly Options
- 6. Navigating Odysseyware Chemistry Test Answers eBook Formats
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 - Interactive Elements Odysseyware Chemistry Test Answers
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Odysseyware Chemistry Test Answers
- 9. Balancing eBooks and Physical Books Odysseyware Chemistry Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Odysseyware Chemistry Test Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Odysseyware Chemistry Test Answers
 - Setting Reading Goals Odysseyware Chemistry Test Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Odysseyware Chemistry Test Answers
 - Fact-Checking eBook Content of Odysseyware Chemistry Test Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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