

CHEMISTRY

2c(ii)

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

where:

$$V_1 = \text{Initial Volume} = 150 \text{ cm}^3$$

$$T_1 = \text{Initial Temp in kelvin} = 21^\circ\text{C} + 273 = 294\text{K}$$

$$T_2 = \text{Final Temp in kelvin} = 45^\circ\text{C} + 273 = 318\text{K}$$

$$V_2 = ?$$

$$V_2 = V_1 \times \frac{T_2}{T_1}$$

$$V_2 = 150 \text{ cm}^3 \times \frac{318\text{K}}{294\text{K}}$$

$$V_2 = 150 \text{ cm}^3 \times 1.0816$$

$$V_2 = 162.24 \text{ cm}^3$$

2d.

I Sodium oxide:



II Phosphorus (III) oxide:



III Chlorine (VII) oxide



IV Magnesium oxide



Real 2014 2015 Waec Chemistry Essay And Objectives

H Kauffman



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