

## List of Common Polyatomic Ions

### Cations:

<u>Symbol</u>	<u>Name</u>	<u>Charge</u>	<u>Symbol</u>	<u>Name</u>	<u>Charge</u>
$\text{NH}_4^+$	Ammonium	+1	$\text{Hg}_2^{+2}$	Dimercury (I)	+2

### Anions:

<u>Symbol</u>	<u>Name</u>	<u>Charge</u>	<u>Symbol</u>	<u>Name</u>	<u>Charge</u>
$\text{C}_2\text{H}_3\text{O}_2^-$	Acetate	-1	$\text{CO}_3^{-2}$	Carbonate	-2
$\text{HCO}_3^-$	Bicarbonate	-1	$\text{CrO}_4^{-2}$	Chromate	-2
$\text{BrO}_3^-$	Bromate	-1	$\text{Cr}_2\text{O}_7^{-2}$	Dichromate	-2
$\text{BrO}_2^-$	Bromite	-1	$\text{C}_2\text{O}_4^{-2}$	Oxalate	-2
$\text{ClO}_4^-$	Perchlorate	-1	$\text{O}_2^{-2}$	Peroxide	-2
$\text{ClO}_3^-$	Chlorate	-1	$\text{SO}_4^{-2}$	Sulfate	-2
$\text{ClO}_2^-$	Chlorite	-1	$\text{SO}_3^{-2}$	Sulfite	-2
$\text{ClO}^-$	Hypochlorite	-1	$\text{S}_2\text{O}_3^{-2}$	Thiosulfate	-2
$\text{CN}^-$	Cyanide	-1	$\text{SiO}_3^{-2}$	Silicate	-2
$\text{SCN}^-$	Thiocyanate	-1	$\text{MoO}_4^{-2}$	Molybdate	-2
$\text{OH}^-$	Hydroxide	-1	$\text{WO}_4^{-2}$	Tungstate	-2
$\text{NO}_3^-$	Nitrate	-1	$\text{BO}_3^{-3}$	Borate	-3
$\text{NO}_2^-$	Nitrite	-1	$\text{PO}_4^{-3}$	Phosphate	-3
$\text{MnO}_4^-$	Permanganate	-1	$\text{PO}_3^{-3}$	Phosphite	-3
$\text{HSO}_4^-$	Bisulfate	-1	$\text{AsO}_4^{-3}$	Arsenate	-3
$\text{HSO}_3^-$	Bisulfite	-1	$\text{AsO}_3^{-3}$	Arsenite	-3
$\text{IO}_3^-$	Iodate	-1			
$\text{IO}_4^-$	Periodate	-1			
$\text{C}_6\text{H}_5\text{COO}^-$	Benzoate	-1			

<u>Symbol</u>	<u>Name</u>	<u>Charge</u>
$\text{H}_2\text{PO}_4^{-1}$	Dihydrogen phosphate	-1
$\text{HPO}_4^{-2}$	Hydrogen phosphate	-2
$\text{Fe}(\text{CN})_6^{-3}$	Hexacyanoferrate (III)	-3
$\text{Fe}(\text{CN})_6^{-4}$	Hexacyanoferrate (II)	-4

# Polyatomic Ions Packet Answers

**Minjie Lin**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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