

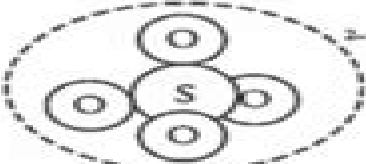
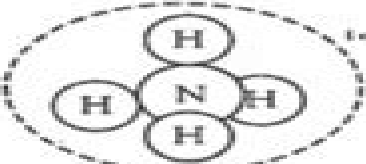
Polyatomic Ions

Can a group of atoms have a charge?

Why?

Do you know you eat a lot of "-ates"? Next time you look at a food label, read the ingredients and you will likely find a number of ingredients that end with "-ate," such as sodium phosphate or calcium carbonate. Did you ever wonder what the chemical formulas of these ingredients look like? In this activity we will explore polyatomic ions, which are groups of atoms that carry a charge. These ions are found in our food ingredients, natural waterways, and many other chemical compounds you encounter every day.

Model 1 – Types of Ions

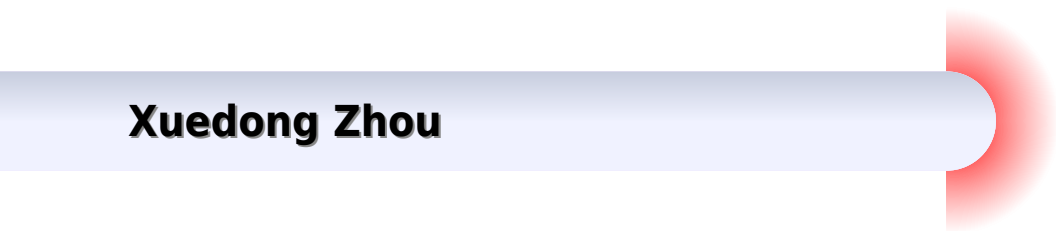
Monatomic Ions	Nitride 	Sulfide 	Chloride 
Polyatomic Ions	Nitrate 	Sulfate 	Ammonium 
	Nitrite 	Sulfite 	Hydroxide 

1. Use Model 1 to complete the table below.

Name of Ion	Nitride	Nitrate	Sulfate	Sulfite	Ammonium
Charge on Ion	-3	-1	-2	-2	+1
Type and Number of Atoms	1 nitrogen	1 nitrogen 3 oxygen	1 sulfur 4 oxygen	1 sulfur 3 oxygen	1 nitrogen 4 hydrogen
Chemical Formula	N^{3-}	NO_3^{1-}	SO_4^{2-}	SO_3^{2-}	NH_4^{1+}

Polyatomic Ions Pogil Activities For High School Chemistry

Xuedong Zhou



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Polyatomic Ions for Surface Analysis and Modification Erick Ryan Fuoco, 2003 **Unimolecular and**
Collision-induced Dissociation Study of Polyatomic Ions at High Collision Energy Xuedong Zhou, 2001 **Quick**
Practice Reproducibles for Chemistry I Tim McDowell, 2012-05-18 A collection of reproducible practice for high school chemistry and middle school physical science Topics include Metric Units Chemical Symbols for elements Atomic Numbers Isotopes Chemical Formulas and Polyatomic Ions **The Particulate Nature of Polyatomic Ions**, 2007 Investigations
Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry Sally M. McIntyre, 2010 *Standard*
Thermodynamic Functions of Gaseous Polyatomic Ions at 100-1000 K Aharon Loewenschuss, Y. Marcus, 1987

Fragmentation of Diatomic and Polyatomic Ions in the Gas Phase Christopher John Proctor, 1981

Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry, 2010 An inductively coupled plasma mass spectrometer ICP MS is an elemental analytical instrument capable of determining nearly all elements in the periodic table at limits of detection in the parts per quadrillion and with a linear analytical range over 8 10 orders of magnitude Three concentric quartz tubes make up the plasma torch Argon gas is spiraled through the outer tube and generates the plasma powered by a looped load coil operating at 27 1 or 40 6 MHz The argon flow of the middle channel is used to keep the plasma above the innermost tube through which solid or aqueous sample is carried in a third argon stream A sample is progressively desolvated atomized and ionized The torch is operated at atmospheric pressure To reach the reduced pressures of mass spectrometers ions are extracted through a series of two approximately one millimeter wide circular apertures set in water cooled metal cones The space between the cones is evacuated to approximately one torr The space behind the second cone is pumped down to or near to the pressure needed for the mass spectrometer MS The first cone called the sampler is placed directly in the plasma plume and its position is adjusted to the point where atomic ions are most abundant The hot plasma gas expands through the sampler orifice and in this expansion is placed the second cone called the skimmer After the skimmer traditional MS designs are employed id est quadrupoles magnetic sectors time of flight ICP MS is the leading trace element analysis technique One of its weaknesses are polyatomic ions This dissertation has added to the fundamental understanding of some of these polyatomic ions their origins and behavior Although mainly continuing the work of others certain novel approaches have been introduced here Chapter 2 includes the first reported efforts to include high temperature corrections to the partition functions of the polyatomic ions in ICP MS This and other objections to preceeding papers in this area were addressed Errors in the measured Tgas values were found for given errors in the experimental and spectroscopic values The ionization energy of the neutral polyatomic ion was included in calculations to prove the validity of ignoring more complicated equilibria Work was begun on the question of agreement between kinetics of the plasma and interface and the increase and depletion seen in certain polyatomic ions This dissertation was also the first

to report day to day ranges for Tgas values and to use a statistical test to compare different operating conditions This will help guide comparisons of previous and future work Chapter 4 was the first attempt to include the excited electronic state 2 in the partition function of ArO as well as the first to address the different dissociation products of the ground and first electronic levels of ArO Chapter 5 reports an interesting source of memory in ICP MS that could affect mathematical corrections for polyatomic ions For future work on these topics I suggest the following experiments and investigations Clearly not an extensive list they are instead the first topics curiosity brings to mind 1 Measurement of Tgas values when using the flow injection technique of Appendix B It was believed that there was a fundamental difference in the plasma when the auto sampler was used versus a continuous injection Is this reflected in Tgas values 2 The work of Chapter 3 can be expanded and supplemented with more trials new cone materials i e copper stainless steel and more cone geometries Some of this equipment is already present in the laboratory others could be purchased or made 3 Tgas values from Chapter 3 could be correlated with instrument pressures during the experiment Pressures after the skimmer cone were recorded for many days but have yet to be collated with the measured Tgas values 4 *Experimental Studies of Polyatomic Ion Interactions with Clean and Adsorbate Covered Metal Surfaces* Samuel B. Wainhaus, 1997

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Table of Contents Polyatomic Ions Pogil Activities For High School Chemistry

1. Understanding the eBook Polyatomic Ions Pogil Activities For High School Chemistry
 - The Rise of Digital Reading Polyatomic Ions Pogil Activities For High School Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ions Pogil Activities For High School Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polyatomic Ions Pogil Activities For High School Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ions Pogil Activities For High School Chemistry
 - Personalized Recommendations
 - Polyatomic Ions Pogil Activities For High School Chemistry User Reviews and Ratings
 - Polyatomic Ions Pogil Activities For High School Chemistry and Bestseller Lists
5. Accessing Polyatomic Ions Pogil Activities For High School Chemistry Free and Paid eBooks

- Polyatomic Ions Pogil Activities For High School Chemistry Public Domain eBooks
 - Polyatomic Ions Pogil Activities For High School Chemistry eBook Subscription Services
 - Polyatomic Ions Pogil Activities For High School Chemistry Budget-Friendly Options
6. Navigating Polyatomic Ions Pogil Activities For High School Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Polyatomic Ions Pogil Activities For High School Chemistry Compatibility with Devices
 - Polyatomic Ions Pogil Activities For High School Chemistry Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Polyatomic Ions Pogil Activities For High School Chemistry
 - Highlighting and Note-Taking Polyatomic Ions Pogil Activities For High School Chemistry
 - Interactive Elements Polyatomic Ions Pogil Activities For High School Chemistry
 8. Staying Engaged with Polyatomic Ions Pogil Activities For High School Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polyatomic Ions Pogil Activities For High School Chemistry
 9. Balancing eBooks and Physical Books Polyatomic Ions Pogil Activities For High School Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Polyatomic Ions Pogil Activities For High School Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Polyatomic Ions Pogil Activities For High School Chemistry
 - Setting Reading Goals Polyatomic Ions Pogil Activities For High School Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Polyatomic Ions Pogil Activities For High School Chemistry
 - Fact-Checking eBook Content of Polyatomic Ions Pogil Activities For High School Chemistry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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