



Oxidation-Reduction Reactions

- Electrons can be neither created out of nothing nor destroyed
- In any redox reaction there is an element being reduced and an element being oxidized
- The total increase in the oxidation numbers must equal the total decrease in the oxidation numbers

If an element is reduced

- It gains electrons
- Oxidation number decreases
- The substance is classified as an oxidizing agent

If an element is oxidized

- It loses electrons
- Oxidation number increases
- The substance is classified as a reducing agent

Oxidation And Reduction Key

S Ben Porath



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