Electrochemistry

Review Oxidation-Reduction (redox) reactions, assigning oxidation numbers.

**LEO goes GER

0 +1-1 +2-1 0

Fe(s) + 2HCl(g) → FeCl₂(s) + H₂(g)
e- transferred from Fe to H*

Oxidizing agent (oxidant) = e' taker (gets reduced)

Reducing agent (reductant) = e' giver (gets oxidized)

Half Reactions

0 Fe \rightarrow Fe²⁺ + 2e⁻ Oxidation 0 2H⁺ + 2e⁻ \rightarrow H₂ Reduction

Balancing equations by the half reaction method.

Often, all of the ions present at the beginning of a redox will not be represented, due to the fact that they are spectators and are not important to the redox process.

Also, redox rxns may be accompanied by the terms "in acid solution or acidic environment" or "in base or basic environment."

- ➤ If acid, use H* and H₂O to balance the equation.
- If base, use OH and H2O to balance the equation.

Acidic env. MnO_4 (aq) + C_2O_4 (aq) $\rightarrow Mn^{2*}$ (aq) + CO_2 (g)

Step 1: Write basic half reactions.

 $MnO_4 \rightarrow Mn^{2*}$ $C_2O_4^{2*} \rightarrow CO_2$

- Step 2: Each half rxn is balanced with respect to the
 - a, elements oxidized and reduced and
 - b, other atoms
 - e, charge

Reduction: 8H* + MnO₄ → Mn²* + 4H₂O (atoms now balance

net chg +7 net chg +2 using H* and H₂O)

The e' change is 5.

So: $5e^{\epsilon} + 8H^{+} + MnO_{4}^{-} \rightarrow Mn^{2+} + 4H_{2}O$

Oxidation: $C_2O_4^{2-} \rightarrow 2CO_2$ (atoms balance)

net chg -2 net chg 0

The e- change is 2. So: $C_2O_4^{2c} \rightarrow 2CO_2 + 2e^{-c}$

Step 3: Balance e' in the 2 half rxns by multiplying by correct values to give a common factor.

e lost <u>must</u> = e gained.

$$2 \times [5c^{2} + 8H^{*} + MnO_{4}^{2} \rightarrow Mn^{2} + 4H_{2}O]$$

$$5 \times [C_{2}O_{4}^{2} \rightarrow 2CO_{2} + 2c^{2}]$$

$$10c^{2} + 16H^{*} + 2MnO_{4}^{2} \rightarrow 2Mn^{2} + 8H_{2}O$$

$$\frac{5C_{2}O_{4}^{2} \rightarrow 10CO_{2} + 10c^{2}}{5C_{2}O_{4}^{2} \rightarrow 2Mn^{2} + 10CO_{2} + 8H_{2}O \text{ (e- cancel)}}$$

$$16H^{*} + 2MnO_{4}^{2} + 5C_{2}O_{4}^{2} \rightarrow 2Mn^{2} + 10CO_{2} + 8H_{2}O \text{ (e- cancel)}$$

Oxidation Reduction And Electrochemistry Section Review

Yijin Wang

Oxidation Reduction And Electrochemistry Section Review:

Decoding **Oxidation Reduction And Electrochemistry Section Review**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "Oxidation Reduction And Electrochemistry Section Review," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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