

Formulae of Elements and Molecules

Complete the following table to show the structure and molecular formula of some covalent compounds. Remember: hydrogen and Group 7 elements form 1 covalent bond, oxygen forms 2, nitrogen forms 3, and carbon forms 4.

Name	Structure	Molecular formula
carbon dioxide		
ammonia		NH <sub>3</sub>
butane		
		HCl

Formulae of Ionic Compounds

Complete the sentences using the words in the box. Use each word once.

negatively gain lost ions  
positively electrostatic

Metal atoms \_\_\_\_\_ electrons and become \_\_\_\_\_ charged ions. Non metal atoms \_\_\_\_\_ electrons and become \_\_\_\_\_

charged ions. The compound formed has \_\_\_\_\_ bonding. The ions are held together by strong \_\_\_\_\_ forces.

Complete the table to show the formula and name for the compounds that could be made from each pair of ions. Think carefully: you must balance the charges. Use a periodic table.

Ion	Li <sup>+</sup>	Na <sup>+</sup>	Mg <sup>2+</sup>	Ca <sup>2+</sup>
F <sup>-</sup>	LiF Lithium fluoride		MgF <sub>2</sub> Magnesium fluoride	
O <sup>2-</sup>				
Br <sup>-</sup>				
Cl <sup>-</sup>				
S <sup>2-</sup>				

C3.1 - Chemical Reactions - Combined Chemistry Revision Sheet

Conservation of Mass

What is the Law of Conservation of mass?

What does it mean when we talk about a closed system?

If a system is not closed (open) why can the masses of substances appear to reduce during an experiment?

If a system is not closed (open) why can the masses of substances appear to increase during an experiment?

Calcium carbonate, CaCO<sub>3</sub>, decomposes when heated. Calcium oxide, CaO, and carbon dioxide, CO<sub>2</sub>, are made. Look at the equation for the reaction.



In an experiment, 20.0g of CaCO<sub>3</sub> are used. 11.2g of CaO are made. What is the mass of CO<sub>2</sub> made?  g

Calculate the relative formula mass of calcium carbonate and calcium oxide (A<sub>r</sub> values: C 12, O 16, Ca 40.1).

Calculate the mass of calcium oxide produced (to the nearest gram) when 10 g of calcium carbonate is heated.

Calculate the mass of carbon dioxide produced (to the nearest gram) when 200 g of calcium carbonate is heated.

Half Equations and Ionic Equations

Complete the sentences using the words in the box. Use each word once.

lost aqueous one don't covalent  
ions water electrons both

Half equations model the change that happens to \_\_\_\_\_ reactant in a chemical reaction and shows how many \_\_\_\_\_ the atoms gain or lose.

The balanced symbol equation for the reaction between sodium and chlorine is:



Write the two half equations for the sodium and chlorine ions and their electrons. Na  + \_\_\_\_\_



An ionic equation shows the \_\_\_\_\_ present in a reaction mixture. Soluble ionic compounds can be identified

in a solution as they have the \_\_\_\_\_

state symbol next to them. Ionic equations also include substances which are simple \_\_\_\_\_ bonded molecules, e.g. \_\_\_\_\_ Look at the following reaction

Identify the ions and write an ionic equation for the reaction.



Spectator ions can be identified on both sides of the equation in the same form and do not change or take part in the reaction. Highlight them in the above equation. Write the net ionic equation for the ions that do change form.

Chemical Equations

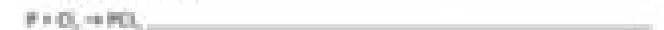
Look at the following equations are they balanced or not balanced?

Count the number of each atom on either side. Write a 0 next to the ones which are balanced and fill next to the ones which are not balanced.

- ☐ Mg + O<sub>2</sub> → MgO
- ☐ CH<sub>4</sub> + Cl<sub>2</sub> → CH<sub>3</sub>Cl + HCl
- ☐ H<sub>2</sub>O<sub>2</sub> → H<sub>2</sub>O + O<sub>2</sub>
- ☐ CH<sub>4</sub> + 2O<sub>2</sub> → CO<sub>2</sub> + 2H<sub>2</sub>O

Chemical equation need to be balanced with the same number of each atoms going into the reaction as reactants as coming out of the reaction as products.

Balance the following equations.



When writing the chemical equations, we

also add state symbols to show the state that a reactant or product is.

State symbol	Meaning
(s)	solid
(l)	liquid
(g)	gas
(aq)	aqueous

# Ocr Gateway Science C3

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