

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the word.

Column I

1. The constant relating unit temperature changes to changes in the volume of a fixed amount of gas at constant pressure
2. The volume of a fixed amount of gas at constant pressure when the absolute temperature is 273 K
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used
4. An experimental gas constant determined from the ideal gas law at 0°C
5. A constant of nature that has values of 8.314 J/mol·K or 0.0821 L·atm/mol·K
6. The a parameter of gas constant temperatures for determining the molar masses of gases
7. An experimental gas constant derived from the ideal gas law at 0°C
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C
10. A unit of energy, the joule

Column II

- a. standard gas law
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. standard molar volume
- h. volume
- i. compressions
- j. heat capacity of oxygen

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Expanding one volume of a monatomic gas from 10 L to 20 L at constant temperature requires _____ J of energy.
a. 10 J
b. 100 J
c. 1000 J
d. 10000 J

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IM Harris



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