The new penny is lighter because it is mostly zinc which has a lower density than copper.

You're The Chemist

- at 90°, mass of 1 drop: 0.019 g at 45°, mass of 1 drop: 0.0218 g at 0°, mass of 1 drop: 0.0242 g Pipets give different results.
- The best angle is 90° because the pipet is easiest to control. Expel the air bubble so that the first drop will be the same size as the rest.
- Find mass of can and divide by density of aluminum. Sample answer: mass of one can: 14.77 g;
- (1) Measure the mass before and after you fill the can with water. Use the mass and density of water to find the volume.
 - (2) Measure the height and radius and calculate volume.
 V = πr²k
 (Can is not a perfect cylinder.)
 - (3) Read label: 12 oz = 355 mL

 $d_{\rm a}$: 2.70 g/cm³: V = 5.47 cm³

- 5. Sample answer:
 - $V = 16.5 \text{ m} \times 3.0 \text{ m} \times 12.8 \text{ m}$ = 630 m³ × 1000 L/m³
 - -6300001J

Assume 30 people with an average weight of 130 lb (1 kg = 2.2 lb) and a density of about 1.0 kg/L.

Volume of 30 people

 $V = 30 \times 130 \text{ lb} \times 1 \text{ kg/2.2 lb} \times 1 \text{ L/1.0 kg}$ = 1800 L

The volume of 30 chairs, 15 tables, and 2 desks is about that of 30 people or 1800 L. The volume of people and furniture is 3600 L. % error = (3600 L/630,000 L)(100%) = 0.57%.

6. If die measures 1.55 cm on a side:

 $V = (1.55 \text{ cm})^3 = 3.72 \text{ cm}^3$

A die has 21 holes that are hemispheres with a radius of 0.20 cm.

Volume of hemisphere: $2/3\pi r^3$ $V = 2/3\pi r^3 = 0.017 \text{ cm}^3$

Volume of 21 hemispheres: 0.36 cm³ Volume of die: $V = 3.72 \text{ cm}^3 - 0.36 \text{ cm}^3$ = 3.36 cm³ Error = 0.36 cm³ Percent error = (0.36 cm³/3.36 cm³)(100%) = 11% Note: The holes in some dice are cones. The volume of a cone is 1/3 π r2h.

 Find weight in pounds and convert to kg. Assume the density is about 1.00 kg/L.
 V = weight × 1 kg/2.2 lb × 1L/1.00 kg

Section Review 4.1

Part A Completion

1. atoms 3. compounds
2. different 4. separated, joined, or rearranged

Part B True-False

5. NT 6. ST 7. AT

Part C Matching

8. c 10. d 9. a 11. b

Part D Questions and Problems

12. simple whole-number ratios

13. $\frac{1 \times 10^6 \text{ atoms}}{\text{gen}} \times \frac{100 \text{ grh}}{\text{m}} = 10^{10} \text{ atoms/m}.$

Section Review 4.2

Part A Completion

1. subatomic S. nucleus
2. negative 6. Ernest Rutherford
3. protons 7. positive
4. neutrons 8. electrons

Part B True-False

9. NT 11. NT 10. AT 12. AT

Part C Matching

13. c 15. e 17. b

Part D Questions and Problems

18. protons, neutrons

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Scott C. Dulebohn

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