

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the item.

Column I

1. The constant relating unit temperature change to the change in volume of a fixed amount of gas at constant pressure
2. The volume of a fixed amount of gas at constant pressure as the absolute temperature is changed
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used
4. An experimental gas constant determined from the ideal gas law at low pressures
5. A constant of nature that has values of about 8.314 J/mol·K
6. The a parameter of gas constant equations for determining the molar mass of a gas
7. An experimental gas constant determined from the ideal gas law
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C
10. A unit of energy, the joule

Column II

- a. molar gas law
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. universal gas constant, pressure
- h. volume
- i. compressibility
- j. definition of joule

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. According to Graham's law, the rate of effusion of a gas is proportional to the square root of its molar mass. Which of the following gases effuses most rapidly?
a. the slowest of the four
b. the one having the least mass
c. the one having the most mass
d. the slowest of the four

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Michael Rajnik



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