

SALT ANALYSIS



SCHEME OF ANALYSIS

| EXPERIMENT | OBSERVATION | INFERENCE |
|---|--|---|
| PRELIMINARY EXAMINATIONS : | | |
| 1) Noted the colour | Colourless | Absence of $\text{Cu}^{2+}, \text{Fe}^{2+}, \text{Fe}^{3+}, \text{Mn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$ |
| | Blue | Pre. of Cu^{2+} |
| | Brown | Pre. of Fe^{3+} |
| | Green | Pre. of Ni^{2+} |
| | Pale pink | Pre. of Mn^{2+} |
| | Pink | Pre. of Co^{2+} |
| 2) Noted the solubility a) In water b) In dil.HCl | a) Soluble and no ppt with Na_2CO_3 b) Insoluble | a) Pre. of NH_4^+ b) Pre. of Pb^{2+} |
| 3) Flame Test : A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame. | a) Blue surrounded by green b) Brick red c) Crimson red d) Bluish white e) Grassy green on prolonged heating | a) Pre. of Cu^{2+} b) Pre. of Ca^{2+} c) Pre. of Sr^{2+} d) Pre. of Pb^{2+} e) Pre. of Ba^{2+} |
| 4) Ash Test : A little of the salt is heated with con. HNO_3 and cobalt nitrate solution. A filter paper is dipped in it and burnt. | a) Green ash b) Blue tinted ash c) Pale pink ash | a) Pre. of Zn^{2+} b) Pre. of Al^{3+} c) Pre. of Mg^{2+} |
| SYSTEMATIC ANALYSIS OF ANIONS : | | |
| 1) $\text{dil.H}_2\text{SO}_4$ test : To the salt add $\text{dil.H}_2\text{SO}_4$, warm if required. | a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky. b) Smell of vinegar | a) Pre. of carbonates b) Pre. of acetate |

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