

Chapter 7 Chemical Reactions

6. What is a double-replacement reaction? A double-replacement reaction is a reaction in which two different compounds exchange positive ions and form two new compounds.
7. Complete the chart by filling in the general forms of the reactions shown.

General Forms	
Single-Replacement Reaction	Double-Replacement Reaction
<u>$A + BC \longrightarrow A + AC$</u>	<u>$AB + CD \longrightarrow AD + CB$</u>

8. Lead(II) nitrate reacts with potassium iodide to form lead(II) iodide and potassium nitrate. Write the balanced equation for this double-replacement reaction. $Pb(NO_3)_2 + 2KI \longrightarrow PbI_2 + 2KNO_3$
9. Circle the letter of the correct answer. Calcium carbonate, $CaCO_3$, reacts with hydrochloric acid, HCl , in a double-replacement reaction. What are the products of this reaction?
- ☒ a. calcium chloride, $CaCl_2$, and carbonic acid, H_2CO_3
 b. calcium hydride, CaH_2 , chlorine, Cl_2 , and carbon dioxide, CO_2
 c. calcium hydrogen carbonate, $Ca(HCO_3)_2$, and chlorine, Cl_2
 d. calcium perchlorate, $Ca(ClO_4)_2$, and methane, CH_4
10. Is the following sentence true or false? A combustion reaction is a reaction in which a substance reacts with carbon dioxide, often producing heat and light. false
11. Methane, CH_4 , burns in oxygen to form carbon dioxide and water. Write the balanced equation for this reaction. $CH_4 + 2O_2 \longrightarrow CO_2 + 2H_2O$
12. Is the following sentence true or false? The reaction that forms water can be classified as either a synthesis reaction or a combustion reaction. true

Reactions as Electron Transfers (pages 204–205)

13. What is an oxidation-reduction reaction? An oxidation-reduction reaction is a reaction in which electrons are transferred from one reactant to another.
14. Calcium reacts with oxygen to form calcium oxide. Which reactant is oxidized in this reaction? calcium
15. Is the following sentence true or false? When calcium reacts with oxygen, each calcium atom gains two electrons and becomes a calcium ion with a charge of $2-$. false
16. Is the following sentence true or false? Oxygen must be present in order for an oxidation-reduction reaction to take place. false
17. The process in which an element gains electrons during a chemical reaction is called reduction.

Prentice Hall Chemistry Hydrocarbon Answers

Robert R. Redfield



Prentice Hall Chemistry Hydrocarbon Answers:

Prentice Hall Chemistry ,2000 **Introductory Organic Chemistry and Hydrocarbons** Caio Lima

Firme,2019-08-28 A novel proposal for teaching organic chemistry based on a broader and simplified use of quantum chemistry theories and notions of some statistical thermodynamic concepts aiming to enrich the learning process of the organic molecular properties and organic reactions A detailed physical chemistry approach to teach organic chemistry for undergraduate students is the main aim of this book A secondary objective is to familiarize undergraduate students with computational chemistry since most of illustrations of optimized geometries plus some topological graphs and information is from quantum chemistry outputs which will also enable students to obtain a deeper understanding of organic chemistry

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11

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functional theory DFT is emphasized and attention is given to the increasing application of artificial intelligence in computational chemistry The book maintains key features from the previous edition delving into the mathematical intricacies of ab initio and density functional methods at an introductory level Clear explanations of matrix methods are provided offering a direct approach to obtaining energy levels and molecular orbitals Additionally each chapter includes sets of Easier and Harder drill questions with suggested answers at the end of the book enhancing the learning experience The book is intended for upper year undergraduate and graduate students studying computational and theoretical chemistry and for self study by researchers in universities and industry to whom computational chemistry may be useful

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