## 11-2 Practice Problems

 Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:

Eglin x 1 mai Liju x 3 mai Liot x 24 g Liot = 0.78 g

 What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?

029 NoI + Cl2 - 2NoCI + Iz 029 NoI + Cl2 - 2NoCI + Iz 029 NoI x 1mol x 2mol NoI (mol NoC) NoCI

> Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:

0+85gCyHo x 1mol x 5mol Cox x 44g = 2.58g 1 x 58g x 2mol CyHo 1mol CO z

 Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:

0.4695bD= x [mol x 2mol 5b, 03 x 121.769 fa.389 Sb

5. What mass of hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>)
must decompose to produce 0.77 g of
water?
2H<sub>2</sub>O<sub>2</sub> -> 2H<sub>2</sub>O + O<sub>2</sub>

0.779 H20 x (mol x 2420 x 349 = [.4594:02

6. What mass of carbon monoxide must react with oxygen to produce 0.69 g of carbon dioxide?

0.69gCOz × 1mol × 2mol CO x 289 = 0.449 1 449 2mol COz 1mol = 0.449  Determine the mass of sodium nitrate produced when 0.73 g of nickel(II) nitrate reacts with sodium hydroxide according to the following equation:

N8(NO<sub>3</sub>)<sub>2</sub> + 2NaOH → N8(OH)<sub>2</sub> + 2NaNO<sub>3</sub> O.73g V8(O<sub>3</sub>)<sub>2</sub> × 1mot × 2mal NaNO<sub>3</sub> × 85g (268g I × 185g × (mol Ni (NO<sub>3</sub>)<sub>2</sub> 1mal NaNO<sub>3</sub>

 Determine the mass of calcium hydroxide produced when calcium carbide reacts with 0.64 g of water according to the following equation:

 $0.64_9H_2O \times \frac{CaC_2 + 2H_2O - Ca(OH)_2 + C_2H_2}{189} \times \frac{1001(a(0H)_2 + C_2H_2)}{1001} \times \frac{749}{1001} = \frac{1.329}{Ca(0H)_2}$ 

How many grams of ozone (O<sub>3</sub>) must decompose to produce 0.87 g of oxygen?

0.8790 x 1mg/ x 2mg/ x 48903 = 0.87909

 Find the mass of sugar (C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>) required to produce 1.82 L of carbon dioxide gas at STP from the reaction described by the following equation: C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> → 2C<sub>2</sub>H<sub>6</sub>O + 2CO<sub>2</sub>

 How many liters of oxygen are necessary for the combustion of 425 g of sulfur, assuming that the reaction occurs at STP? The balanced equation is S + O<sub>2</sub> → SO<sub>2</sub>.

 Find the mass of benzene (C<sub>6</sub>H<sub>6</sub>) required to produce 2.66 L of carbon dioxide gas at STP from the reaction described by the following equation:

2C4H6 + 15O2 → 6H2O + 12CO2

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**Rachel S Tattersall** 

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