

7-3 Review and Reinforcement (continued)

Write the name of each of the following ionic compounds.

- | | | |
|--------------------------|----------------------------|-------|
| 17. ZnO _____ | zinc oxide | _____ |
| 18. $NaNO_3$ _____ | sodium nitrate | _____ |
| 19. K_2CO_3 _____ | potassium carbonate | _____ |
| 20. $Mg(OH)_2$ _____ | magnesium hydroxide | _____ |
| 21. K_2S _____ | potassium sulfide | _____ |
| 22. CaF _____ | calcium fluoride | _____ |
| 23. $AgCl$ _____ | silver chloride | _____ |
| 24. $(NH_4)_3PO_4$ _____ | ammonium phosphate | _____ |

Write the name of each of the following molecular compounds.

- | | | |
|--------------------|---------------------------------|-------|
| 25. SO_2 _____ | sulfur dioxide | _____ |
| 26. HI _____ | hydroiodic acid | _____ |
| 27. P_2S_5 _____ | diphosphorus trisulfide | _____ |
| 28. NF_3 _____ | nitrogen trifluoride | _____ |
| 29. SF_6 _____ | sulfur hexafluoride | _____ |
| 30. H_2S _____ | hydrosulfuric acid | _____ |
| 31. N_2O_5 _____ | dinitrogen pentoxide | _____ |
| 32. PCl_5 _____ | phosphorus pentachloride | _____ |

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Chemistry of +1 Iodine in Alkaline Solution Yuan-tsan Chia, 1958 [Ionic Soft Matter: Modern Trends in Theory and Applications](#) Douglas Henderson, Myroslav Holovko, Andrij Trokhymchuk, [Solution Thermodynamics and its Application to Aqueous Solutions](#) Yoshikata Koga, 2007-11-12 As the title suggests we introduce a novel differential approach to solution thermodynamics and use it for the study of aqueous solutions We evaluate the quantities of higher order derivative than the normal thermodynamic functions We allow these higher derivative data speak for themselves without resorting to any model system We thus elucidate the molecular processes in solution referred to in this book mixing scheme to the depth equal to if not deeper than that gained by spectroscopic and other methods We show that there are three composition regions in aqueous solutions of non electrolytes each of which has a qualitatively distinct mixing scheme The boundary between the adjacent

regions is associated with an anomaly in the third derivatives of G . The loci of the anomalies in the temperature composition field form the line sometimes referred to as Koga line. We then take advantage of the anomaly of a third derivative quantity of 1-propanol in the ternary aqueous solution 1-propanol sample species H_2O . We use its induced change as a probe of the effect of a sample species on H_2O . In this way we clarified what a hydrophobe or a hydrophile and in turn an amphiphile does to H_2O . We also apply the same methodology to ions that have been ranked by the Hofmeister series. We show that the kosmotropes salting out or stabilizing agents are either hydrophobes or hydration centres and that chaotropes salting in or destabilizing agents are hydrophiles. A new differential approach to solution thermodynamics. A particularly clear elucidation of the mixing schemes in aqueous solutions. A clear understanding on the effects of hydrophobes, hydrophiles and amphiphiles to H_2O . A clear understanding on the effects of ions on H_2O in relation to the Hofmeister effect. A new differential approach to studies in multi-component aqueous solutions.

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