

Chemistry Chapter 8 Test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ____ 1. What type of bond would you expect in a compound of oxygen and sodium?
 (a) ionic c. covalent
 b. metallic d. polar covalent
- ____ 2. The symbols δ^- and δ^+ indicate which of the following?
 a. electronegativity c. charge
 b. electrostatic force (d) partial charge
- ____ 3. Which of the following represents a nonpolar covalent molecule?
 a. NaCl c. HCN
 (b) Br₂ d. Au
- ____ 4. The measure of an atom's ability to attract electrons to itself in a chemical bond is called _____.
 a. electron affinity (c) electronegativity
 b. polarity d. ionization energy
- ____ 5. The polarity of a bond depends on the difference in the atoms' _____.
 a. electron affinities c. atomic radii
 (b) electronegativities d. ionization energies
- ____ 6. What type of bond would you expect to form between nitrogen and potassium?
 (a) ionic c. polar covalent
 b. covalent d. metallic
- ____ * High electronegativity differences are found in _____ bonds.
 a. metallic c. covalent
 (b) ionic metal - nonmetal d. Pauling
- ____ * 8. To achieve an octet when combining with potassium, would an atom of a Group 16 element be expected to lose, gain, or share electrons?
 a. lose K
 (b) gain K 17
 c. share
 (d) This cannot be determined from the given information. x
- ____ 9. Which of the following does not exist as a diatomic molecule?
 (a) argon c. chlorine
 b. hydrogen d. oxygen
 H N O F Cl Br I
 Gas: H₂ N₂ O₂ F₂ Cl₂ Br₂ I₂
- ____ 10. Which of these is not a covalent molecule?
 a. F₂ c. LiF
 b. CO₂ d. NH₃
- ____ 11. A dash in a Lewis structure represents _____.
 a. one electron c. four electrons
 (b) two electrons d. unpaired electrons
- ____ 12. The central atom in a Lewis structure is usually the atom _____.

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incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science centers museums and zoos where teachers can take middle school students for interactive science experiences Another chapter describes nearly 140 professional associations and U S government agencies that offer resources and assistance Authoritative extensive and thoroughly indexed and the only guide of its kind *Resources for Teaching Middle School Science* will be the most used book on the shelf for science teachers school administrators teacher trainers science curriculum specialists advocates of hands on science teaching and concerned parents *Fundamentals of Chemistry* Burns A.

Ralph, Charles Baker, Claire Baker, 1995 This book is a supplement to the texts not a replacement It is intended to maximize your success in this course by showing you how to become involved in developing your own techniques for grasping the concepts of chemistry Using a study outline problem sets problem examples worked and unworked and numerous self tests with answers this manual will provide you with opportunities to sharpen your skills and evaluate your comprehension of the material in the texts The worked out solutions at the end of this manual walk you step by step through the methods of arriving at the answer to those same problems which have an answer only in the answer key Used in conjunction with your textbook and classroom lecture notes this book offers an essential learning opportunity to the chemistry student Back cover

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