

## 11-1 Practice Problems

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1. Lead will react with Hydrochloric acid to produce Lead (II) Chloride and Hydrogen gas. How many moles of Hydrochloric acid are needed to completely react with 0.36-mol of lead?

2. How many moles of  $\text{HNO}_3$  will be produced when 0.51-mol of  $\text{N}_2\text{O}_5$  reacts according to the following equation?  
 $\text{N}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3$

3. Iron will react with Oxygen Gas to produce  $\text{Fe}_2\text{O}_3$ . How many moles of  $\text{Fe}_2\text{O}_3$  will be produced if 0.18-mol of Fe reacts?

4. How many moles of NaBr will be produced when 0.69-mol of Bromine reacts according to the following equation?  
 $\text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + \text{I}_2$

5. Phosphorus will react with Bromine to produce Phosphorus Tribromide. How many moles of Phosphorus Tribromide will be produced if 0.78-mol of Bromine is reacted?

6. Carbon will react with Zinc Oxide to produce Zinc and Carbon Dioxide. How many moles of carbon dioxide will be produced if 0.38-mol of  $\text{ZnO}$  is completely reacted?

7. How many moles of oxygen will be needed to react with 0.38-mol of  $\text{C}_3\text{H}_8$  according to the following equation?  
 $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

8. Nitrogen gas can react with Hydrogen gas to produce ammonia. How many moles of Nitrogen will be needed to produce 0.48-mol of  $\text{NH}_3$ ?

9. How many moles of hydrogen will be produced if 0.44-mol of  $\text{CaH}_2$  reacts according to the following equation?  
 $\text{CaH}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + 2\text{H}_2$

10. How many moles of water will be produced if 2.35-mol of oxygen reacts according to the following equation?  
 $2\text{C}_6\text{H}_6 + 15\text{O}_2 \rightarrow 12\text{CO}_2 + 6\text{H}_2\text{O}$

# Prentice Hall Chemistry Answer Key Practice Problems

**Guillaume Favre**



**Prentice Hall Chemistry Answer Key Practice Problems:**

The book delves into Prentice Hall Chemistry Answer Key Practice Problems. Prentice Hall Chemistry Answer Key Practice Problems is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Prentice Hall Chemistry Answer Key Practice Problems, encompassing both the fundamentals and more intricate discussions.

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    - Chapter 5: Conclusion
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  3. In chapter 2, the author will delve into the foundational concepts of Prentice Hall Chemistry Answer Key Practice Problems. The second chapter will elucidate the essential principles that must be understood to grasp Prentice Hall Chemistry Answer Key Practice Problems in its entirety.
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  6. In chapter 5, the author will draw a conclusion about Prentice Hall Chemistry Answer Key Practice Problems. The final chapter will summarize the key points that have been discussed throughout the book.
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