

## Chapter 7: Chemical Reaction Guided Notes

### I. Chemical Changes in Matter

#### A. Chemical Reaction

1. A change in which one or more substances are converted to \_\_\_\_\_ substances.

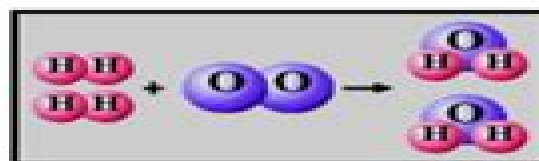


#### B. Law of Conservation of Matter

1. In a chemical reaction, matter is not \_\_\_\_\_ or \_\_\_\_\_.
2. Atoms can only be \_\_\_\_\_.
3. Discovered by Lavoisier.

\_\_\_\_\_ Hydrogen

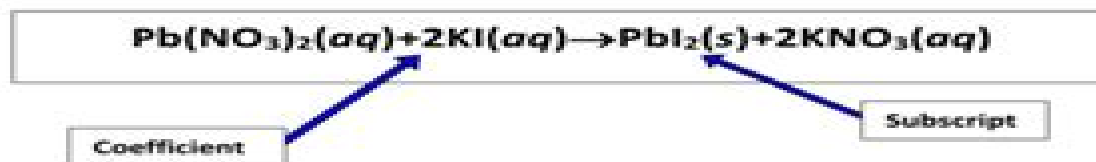
\_\_\_\_\_ Oxygen



\_\_\_\_\_ Hydrogen

\_\_\_\_\_ Oxygen

#### C. Chemical Equations



1. \_\_\_\_\_ - # of units of each substance.
2. Individual atom = **atom**  
 $2\text{Mg} \Rightarrow 2$  \_\_\_\_\_ of magnesium
3. Covalent substance = **molecule**  
 $3\text{CO}_2 \Rightarrow 3$  \_\_\_\_\_ of carbon dioxide
4. Ionic substance = **unit**  
 $4\text{MgO} \Rightarrow 4$  \_\_\_\_\_ of magnesium oxide

| SYMBOL                 | MEANING                            |
|------------------------|------------------------------------|
| $\longrightarrow$      | produces, forms                    |
| +                      | plus, and                          |
| (s)                    | solid                              |
| (l)                    | liquid                             |
| (g)                    | gas                                |
| (aq)                   | aqueous (solid dissolved in water) |
| $\xrightarrow{\Delta}$ | the reactants are heated           |

# Science Chemical Reactions Note Taking Guide Answers

**S Ashworth**



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