

Chapter 3 Review Questions

Question: 2, 4, 12 (d), 12 (e), 16 (d), 16 (e), 18 (d), 18 (e), 26, 34 (a), 34 (c), 38, 40

2. What is the total mass of hydrogen in each of the molecules? *atomic mass (H) = 1.008 g/mol*
- (a) CH_4
 - (b) CHCl_3
 - (c) $\text{C}_{12}\text{H}_{10}\text{O}_6$
 - (d) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

a) $\text{H} = 4$

Total mass (H) = $4 \times 1.008 \text{ g/mol}$

Total mass (H) = 4.032 amu

a) $\text{H} = 12$

Total mass (H) = $12 \times 1.008 \text{ g/mol}$

Total mass (H) = 12.096 amu

b) $\text{H} = 1$

Total mass (H) = $1 \times 1.008 \text{ g/mol}$

Total mass (H) = 1.008 amu

c) $\text{H} = 10$

Total mass (H) = $10 \times 1.008 \text{ g/mol}$

Total mass (H) = 10.08

4. Determine the molecular mass of the following compounds:

(a)



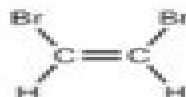
molar mass = $(2 \times 35.45) + (1 \times 12.01) + (1 \times 15.9) = 98.81 \text{ g/mol}$

(b)



molar mass = $(2 \times 1.008) + (2 \times 12.01) = 26.038 \text{ g/mol}$

(c)



molar mass = $(2 \times 1.008) + (2 \times 79.904) + (2 \times 12.01) = 165.84 \text{ g/mol}$

(d)



molar mass = $(2 \times 16.00) + (4 \times 15.999) + (2 \times 1.008) = 98.078 \text{ g/mol}$

Review Questions For Hewitt Chapte 35

**Shelley Cobbett, Anne G. Perry, Patricia
A. Potter, Wendy R. Ostendorf**

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