

Periods and Groups: Valence Electrons: Memory

Periods and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group					Group	
	1A	2A	3A	4A	5A	6A	7A	8A	1A
1	H	He							He
2	Li	Be							Ne
3	Na	Mg							Ar
4	K	Ca							Kr

Thinking like the game "Memory", answer that can element is in period 3 and group 1.

What element is in Period 2, and Group 16A? (Group 16 is also called 6A.)

What element is in Period 3 and Group 1?

What element is in Group 2A, and Period 2?

What group number is in Period 4?

What group number is in Period 3?

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains elements that have the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Lithium 3



3 valence electrons

Hydrogen and Helium are both in Group 1A (also called 1) and have 1 valence electron.

Group 1A, including H and He, has 1 valence electron. How many valence electrons will they have?

How many electrons?

Group 2A: - Shows you more clearly that there is 1 shell of electrons. The elements in Group 2A and 2B are called 2A and 2B. Helium, which is in Group 1A, has 2 valence electrons. "He 2A, 2B" is in Group 1A, but it has 2 valence electrons, so it's in Group 2A. Helium is in Group 1A, but it has 2 valence electrons, so it's in Group 2A.

Periods And Groups Valence Electrons Masses Answers

Lingjun Ying



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