

5 - Periodicity Activity

1 2 3 4 5 6 7 8 9 10 11 12

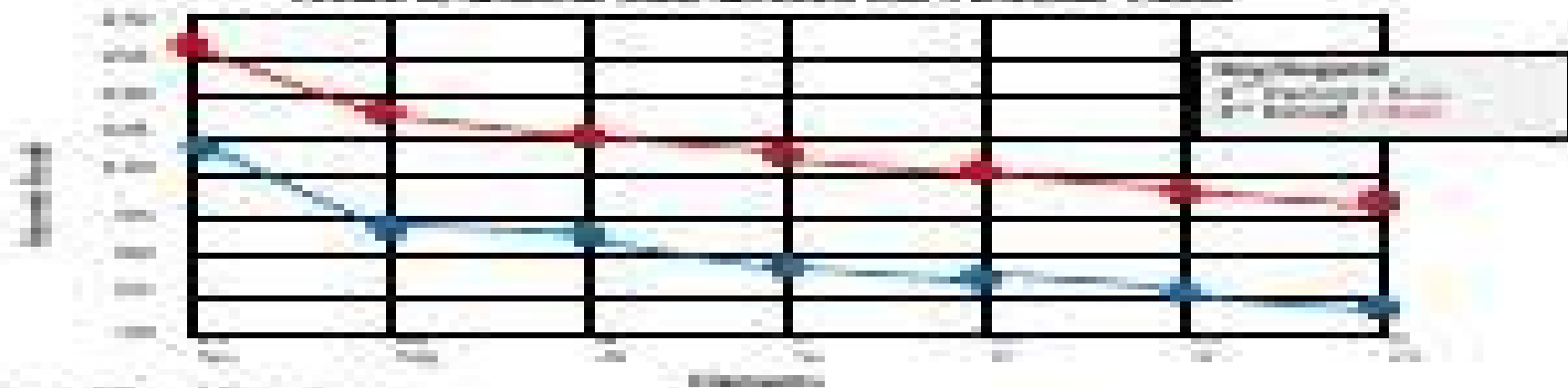
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph below best fits graphing Period 2's data points (graph below)
2. Is a slight downward curve or does graph Period 2's data fit the other graph?
3. Label the line as ionisation enthalpy
4. decreases downwards...

	1	2	3	4	5	6	7	8
1	Li	Be	B	C	N	O	F	Ne
2	Na	Mg	Al	Si	P	S	Cl	Ar
3	K	Ca	Sc	Ti	V	Cr	Mn	Fe
4	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru
5	Cs	Ba	La	Hf	Ta	W	Re	Os
6	Fr	Ra	Ac	Th	Pa	U	Np	Pu

Now label the elements above in the following way:

Trend in Atomic Size across the Periodic Table



Illustrated Trend Equations

1. Following the trend in atomic size across the periodic table decreases
2. This means two atoms that are more metals produce a difference that goes in decreasing spelling differences related to increasing metals

Worked Trend Equations

1. Periodic trend of electronegativity: increases (E.g. Na, K, Rb, Cs, Fr has an increasing movement as you go down in column of the periodic table) increases
2. This means two atoms that are more metals produce a difference (a pair of differences) making the following differences:

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PT Brinkman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the PT Brinkman text.

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