

Periodic Trends: Straw Lab

Objective: To create a visual representation of the following periodic trends: atomic size, ionization energy and electronegativity.

Atomic Radius: the size of an atom measured in either nanometers (nm) or angstroms (Å).

Ionization Energy: the energy needed to remove an electron from an atom measured in kilojoules per mole (kJ/mol) of an atom.

Electronegativity: the relative ability of an atom to attract electrons towards itself.

Materials: For this lab you will need a sheet with 3 small periodic tables, straws, scissors, a metric ruler, your notebook and access to a hot glue gun.

Procedure:

- 1) Draw and complete the data table for each trend in your lab notebook. Each person should have their own data tables in their own notebooks!
- 2) For each trend, calculate the scale factor according to the instructions. Using the scale factor, measure the straw and cut it to the appropriate size.
- 3) You will then use the glue gun to glue the straw to the periodic table in the box that represents the element's location on the periodic table.
- 4) Answer the questions for each trend in your lab notebook. Everyone should have their own answers in their own notebook!

Atomic Radius:

The radius of the following atoms has been estimated to have the following values:

Name	Symbol	Atomic Number	Size of Radius (nm)	Calculated Straw Length (cm)
Fluorine			0.064	
Chlorine			0.0999	
Bromine			0.114	
Iodine			0.1133	
Sodium			0.186	
Magnesium			0.160	
Aluminum			0.143	
Silicon			0.117	

Periodic Trends Straw Lab Answer Key

Michal Rosen-Zvi



Periodic Trends Straw Lab Answer Key:

Decoding **Periodic Trends Straw Lab Answer Key**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Periodic Trends Straw Lab Answer Key**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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Table of Contents Periodic Trends Straw Lab Answer Key

1. Understanding the eBook Periodic Trends Straw Lab Answer Key
 - The Rise of Digital Reading Periodic Trends Straw Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Straw Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Straw Lab Answer Key
 - Personalized Recommendations
 - Periodic Trends Straw Lab Answer Key User Reviews and Ratings
 - Periodic Trends Straw Lab Answer Key and Bestseller Lists

5. Accessing Periodic Trends Straw Lab Answer Key Free and Paid eBooks
 - Periodic Trends Straw Lab Answer Key Public Domain eBooks
 - Periodic Trends Straw Lab Answer Key eBook Subscription Services
 - Periodic Trends Straw Lab Answer Key Budget-Friendly Options
6. Navigating Periodic Trends Straw Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Straw Lab Answer Key Compatibility with Devices
 - Periodic Trends Straw Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Straw Lab Answer Key
 - Highlighting and Note-Taking Periodic Trends Straw Lab Answer Key
 - Interactive Elements Periodic Trends Straw Lab Answer Key
8. Staying Engaged with Periodic Trends Straw Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Straw Lab Answer Key
9. Balancing eBooks and Physical Books Periodic Trends Straw Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Straw Lab Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Straw Lab Answer Key
 - Setting Reading Goals Periodic Trends Straw Lab Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Straw Lab Answer Key
 - Fact-Checking eBook Content of Periodic Trends Straw Lab Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton ha + b → a⁻ + hb strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web acid base titrations laboratory report name rj hamdan date 11/8 section 432 the goal of this experiment is using the skills of precision to determine the

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web jul 19 2023 at the beginning of the titration the solution in the erlenmeyer flask is acidic as the base is added it completely reacts with the acid and the solution in the erlenmeyer flask continues to be acidic but at the equivalence point

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web spreadsheet computations prelab report this is the type of computations the computer does to determine first and second derivatives refer to these spreadsheet columns for the comments below pH titration graph lab data provide columns a and b

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web 1 approximately 3 g of hydrated oxalic acid $H_2C_2O_4 \cdot 2H_2O$ was weighted to the nearest 0.01 g in a 50 mL beaker 2 about

30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution pH is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then $\text{pH} = -\log 0.100 = 1.000$ pH $\log 0.100 = 1.000$ b titrant volume 12.50 ml

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