

# The Periodic Table of Elements

In 1869 Dimitri Mendeleev organized the elements in a table based on the properties of the elements.

Groups

↓ ↓  
(Columns)

The elements in each group have the same number of valence electrons.

|                                       |                                             |                   |                                    |                    |                                    |                   |                                    |                    |                                    |                    |                                    |                   |                                    |                   |                                    |                  |                                               |
|---------------------------------------|---------------------------------------------|-------------------|------------------------------------|--------------------|------------------------------------|-------------------|------------------------------------|--------------------|------------------------------------|--------------------|------------------------------------|-------------------|------------------------------------|-------------------|------------------------------------|------------------|-----------------------------------------------|
| 1 <sup>st</sup> Period → (1 shell) ☉  | Group 1-2<br>(1, 2 valence e <sup>-</sup> ) | 1 H<br>Hydrogen   | Group 2-2<br>(2 valence electrons) | 4 Be<br>Beryllium  | Group 3-2<br>(3 valence electrons) | 5 B<br>Boron      | Group 4-2<br>(4 valence electrons) | 6 C<br>Carbon      | Group 5-2<br>(5 valence electrons) | 7 N<br>Nitrogen    | Group 6-2<br>(6 valence electrons) | 8 O<br>Oxygen     | Group 7-2<br>(7 valence electrons) | 9 F<br>Fluorine   | Group 8-2<br>(8 valence electrons) | 2 He<br>Helium   | ↳ Except Helium with 2 valence e <sup>-</sup> |
| 2 <sup>nd</sup> Period → (2 shells) ☉ |                                             | 3 Li<br>Lithium   |                                    |                    |                                    |                   |                                    |                    |                                    |                    |                                    |                   |                                    |                   |                                    | 10 Ne<br>Neon    |                                               |
| 3 <sup>rd</sup> Period → (3 shells) ☉ |                                             | 11 Na<br>Sodium   |                                    | 12 Mg<br>Magnesium |                                    | 13 Al<br>Aluminum |                                    | 14 Si<br>Silicon   |                                    | 15 P<br>Phosphorus |                                    | 16 S<br>Sulfur    |                                    | 17 Cl<br>Chlorine |                                    | 18 Ar<br>Argon   |                                               |
| 4 <sup>th</sup> Period → (4 shells) ☉ |                                             | 19 K<br>Potassium |                                    | 20 Ca<br>Calcium   |                                    | 31 Ga<br>Gallium  |                                    | 32 Ge<br>Germanium |                                    | 33 As<br>Arsenic   |                                    | 34 Se<br>Selenium |                                    | 35 Br<br>Bromine  |                                    | 36 Kr<br>Krypton |                                               |

Periods →  
(rows)

The elements in each period have the same number of shells.

Today, 114 elements are arranged by atomic number in the periodic table. The atomic number is based on the number of protons in an element.

# Periodic Table Notes For Middle School

**Michal Rosen-Zvi**



## **Periodic Table Notes For Middle School:**

## Unveiling the Power of Verbal Artistry: An Emotional Sojourn through **Periodic Table Notes For Middle School**

In a global inundated with monitors and the cacophony of instantaneous conversation, the profound energy and mental resonance of verbal art often fade into obscurity, eclipsed by the continuous assault of noise and distractions. Yet, located within the lyrical pages of **Periodic Table Notes For Middle School**, a captivating perform of literary splendor that pulses with organic thoughts, lies an wonderful journey waiting to be embarked upon. Penned with a virtuoso wordsmith, this enchanting opus books viewers on a mental odyssey, gently revealing the latent possible and profound affect stuck within the delicate web of language. Within the heart-wrenching expanse of this evocative evaluation, we shall embark upon an introspective exploration of the book is central themes, dissect its captivating publishing fashion, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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