

Periodic Trends

1. Explain the following trends:

- Atomic Radius
- Electronegativity
- Ionization Energy

2. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

3. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

4. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

5. Order the following elements by increasing metallic character:

Li, Na, K, Rb, Cs, Fr

6. Order the following elements by increasing non-metallic character:

F, Cl, Br, I, At

7. Order the following elements by increasing atomic mass:

Li, Na, K, Rb, Cs, Fr

8. Order the following elements by increasing atomic number:

Li, Na, K, Rb, Cs, Fr

9. Order the following elements by increasing atomic weight:

Li, Na, K, Rb, Cs, Fr

10. Order the following elements by increasing atomic weight:

Li, Na, K, Rb, Cs, Fr

Name: _____ Date: _____ Period: _____

Periodic Trends Test Study Guide

Section 1: Atomic Structure and Properties of the Periodic Table

1. What are the three main categories of elements in the periodic table?

- Metals
- Non-metals
- Metalloids

2. List the following elements in order of increasing atomic radius: Li, Na, K, Rb, Cs, Fr.

3. List the following elements in order of increasing electronegativity: F, Cl, Br, I, At.

4. List the following elements in order of increasing ionization energy: Li, Na, K, Rb, Cs, Fr.

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Section 2: Periodic Trends

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2. List the following elements in order of increasing atomic radius: Li, Na, K, Rb, Cs, Fr.

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7. List the following elements in order of increasing atomic mass: Li, Na, K, Rb, Cs, Fr.

8. List the following elements in order of increasing atomic number: Li, Na, K, Rb, Cs, Fr.

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Section 3: Periodic Trends

1. What are the three main categories of elements in the periodic table? (Metals, Non-metals, Metalloids)

2. List the following elements in order of increasing atomic radius: Li, Na, K, Rb, Cs, Fr.

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Mark Steyvers



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