

## Percent Yield Practice

1. Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.



$$\frac{24.8 \text{ g CaCO}_3}{1} \cdot \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \cdot \frac{1 \text{ mol CaO}}{1 \text{ mol CaCO}_3} \cdot \frac{56.08 \text{ g CaO}}{1 \text{ mol CaO}} = 13.9 \text{ g CaO} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{13.1}{13.9} \cdot 100 = \boxed{94.24\%} \Rightarrow \boxed{94.2\%}$$

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?



$$\frac{1.87 \text{ g Al}}{1} \cdot \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol Al}} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 16.6 \text{ g Cu} \quad \text{TY}$$

$$\frac{9.65 \text{ g CuSO}_4}{1} \cdot \frac{1 \text{ mol CuSO}_4}{159.62 \text{ g CuSO}_4} \cdot \frac{3 \text{ mol Cu}}{2 \text{ mol CuSO}_4} \cdot \frac{63.55 \text{ g Cu}}{1 \text{ mol Cu}} = 3.84 \text{ g Cu} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100 = \frac{3.65}{3.84} \cdot 100 = \boxed{95.05\%}$$

3. If you start with 5.0 g of  $\text{C}_3\text{H}_8$  and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?



$$\frac{5.0 \text{ g C}_3\text{H}_8}{1} \cdot \frac{1 \text{ mol C}_3\text{H}_8}{44.11 \text{ g C}_3\text{H}_8} \cdot \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 8.2 \text{ g H}_2\text{O} \quad \text{TY}$$

$$\text{PY} = \frac{\text{AY}}{\text{TY}} \cdot 100$$

$$\frac{\text{PY} \cdot \text{TY}}{100} = \text{AY} = \frac{75 \cdot 8.2}{100} = \boxed{6.2 \text{ g H}_2\text{O}} \quad \text{AY}$$

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?



$$\frac{20 \text{ g CaO}}{1} \cdot \frac{1 \text{ mol CaO}}{56.08 \text{ g CaO}} \cdot \frac{1 \text{ mol Na}_2\text{O}}{1 \text{ mol CaO}} \cdot \frac{61.98 \text{ g Na}_2\text{O}}{1 \text{ mol Na}_2\text{O}} = \boxed{20 \text{ g Na}_2\text{O}} \quad \text{TY}$$

# Percent Yield Guided Problems And Answers

**Yijin Wang**



## **Percent Yield Guided Problems And Answers:**

## **Percent Yield Guided Problems And Answers Book Review: Unveiling the Power of Words**

In a global driven by information and connectivity, the power of words has be evident than ever. They have the capability to inspire, provoke, and ignite change. Such may be the essence of the book **Percent Yield Guided Problems And Answers**, a literary masterpiece that delves deep in to the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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### **Table of Contents Percent Yield Guided Problems And Answers**

1. Understanding the eBook Percent Yield Guided Problems And Answers
  - The Rise of Digital Reading Percent Yield Guided Problems And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Percent Yield Guided Problems And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Percent Yield Guided Problems And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Percent Yield Guided Problems And Answers
  - Personalized Recommendations
  - Percent Yield Guided Problems And Answers User Reviews and Ratings
  - Percent Yield Guided Problems And Answers and Bestseller Lists

5. Accessing Percent Yield Guided Problems And Answers Free and Paid eBooks
  - Percent Yield Guided Problems And Answers Public Domain eBooks
  - Percent Yield Guided Problems And Answers eBook Subscription Services
  - Percent Yield Guided Problems And Answers Budget-Friendly Options
6. Navigating Percent Yield Guided Problems And Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Percent Yield Guided Problems And Answers Compatibility with Devices
  - Percent Yield Guided Problems And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Percent Yield Guided Problems And Answers
  - Highlighting and Note-Taking Percent Yield Guided Problems And Answers
  - Interactive Elements Percent Yield Guided Problems And Answers
8. Staying Engaged with Percent Yield Guided Problems And Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Percent Yield Guided Problems And Answers
9. Balancing eBooks and Physical Books Percent Yield Guided Problems And Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Percent Yield Guided Problems And Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Percent Yield Guided Problems And Answers
  - Setting Reading Goals Percent Yield Guided Problems And Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Percent Yield Guided Problems And Answers
  - Fact-Checking eBook Content of Percent Yield Guided Problems And Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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