

Gay-Lussac's Law

6. The gas in an aerosol can is at a temperature of 27°C and a pressure of 5 atm. If the gas reaches a pressure of 9 atm, what will the can's temperature be?
7. An oxygen tank reads 150 kPa with a temperature of 25°C . What will the pressure read if the tank is cooled to -183°C ?

Combined Gas Law

8. A gas has a volume of 35 liters at STP. What will its volume be at 5.0 atm and 35°C ?
9. 0.4 L of a gas is contained at 129 atm and 273 K. What will the volume of the gas be at 15 atm and 293K?
10. 400 mL of gas is stored at STP. The volume changes to 560 mL, and the temperature to 30°C . What is the pressure (in atm)?

Ideal Gas Law

11. If 6 moles of a gas is at a pressure of 6.6 atm and a volume of 12 liters, what is the temperature of the gas?
12. If there is unknown quantity of gas at a pressure of 132.6 kPa, a volume of 25 liters, and a temperature of 87°C , how many moles of gas are in the sample?
13. If there are 5 moles of gas in a 50-liter container, at a temperature of 250 K, what is the pressure of the gas within the container?
14. If 10.5 moles of gas are stored at a pressure of 0.15 atm and a temperature of 25°C , what is the volume of the container that the gas is stored in?

Dalton's Law

15. A storage container holds a mixture of oxygen, helium, and carbon dioxide. If the partial pressures of the three gases tank are 15 atm of O_2 , 8 atm of He, and 25 atm of CO_2 , what is the total pressure inside of the container?
16. The total pressure of a gas mixture is 1.9 kPa. The partial pressure of gas A is 0.06 kPa and the partial pressure of gas B is 0.25 kPa. What is the partial pressure of the remaining gas in the mixture?
17. If I place 50 moles of N_2 and 1.5 moles of CO_2 in a 40 L container at a temperature of 25°C , what will the pressure of the resulting mixture of gases be?

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**David I. Ketcheson,Randall J.
LeVeque,Mauricio J. del Razo**



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