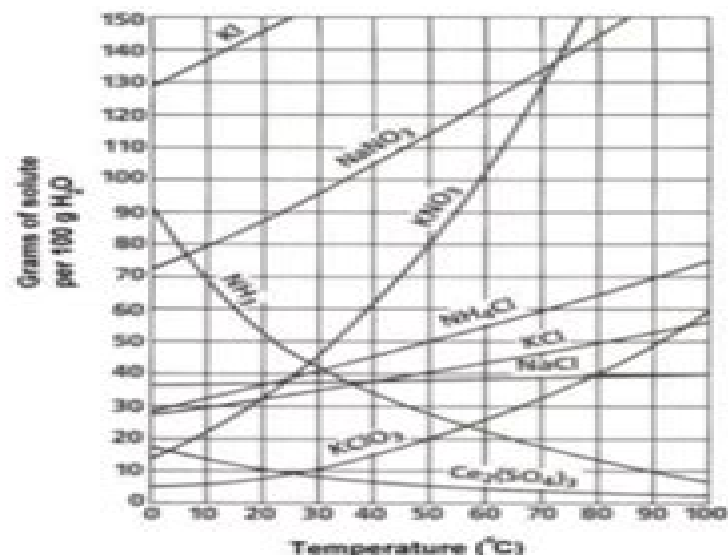


Chemistry I-H
Solubility Curves and Solutions Review Sheet
Solution Set

1. Define solubility: **The maximum amount of solute that dissolves in a fixed quantity of a particular solvent at a specified temperature.**

Use the solubility curve below to answer questions #2 - #15. Be sure to note the units on the axes of the graph.



2. In general, how does temperature affect solubility?
The higher the temp. the more substance dissolves
3. **KClO₃** Which compound is least soluble at 10°C?
4. **50 g** How many grams of KCl can be dissolved in 100g of water at 80°C?
5. **40 g** How many grams of NaCl can be dissolved in 100g of water at 90°C?
6. **about 60 g** At 40°C, how much KNO₃ can be dissolved in 100g of water?
7. **NaCl** Which compound shows the least amount of change in solubility from 0°C-100°C?
8. **unsaturated** At 30°C, 90g of NaNO₃ is dissolved in 100g of water. Is this solution saturated or unsaturated?
9. **D** At 60 C, 72 g of NH₄Cl are dissolved in 100g of water. This solution is considered to be:
A) saturated; B) unsaturated; C) supersaturated; D) saturated with some left undissolved
10. **30 g** A saturated solution of KClO₃ is formed from one hundred grams of water. If the saturated solution is cooled from 90°C to 50°C, how many grams of precipitate are formed?
11. **20 g** A saturated solution of NH₄Cl is formed from one hundred grams of water. If the saturated solution is cooled from 80°C to 40°C, how many grams of precipitate are formed?
12. **E** Which compounds show a *decrease* in solubility from 0°C-100°C?
A) potassium chlorate B) cerium sulfate C) sodium nitrate D) ammonia E) Both B and D
13. **C** Which compound is the most soluble at 10°C?
A) Ce₂(SO₄)₃ B) NaCl C) KCl D) KClO₃ E) NH₃
14. **C** A chemist adds 120 grams of KCl into a test tube which contains 100 grams of water. The temperature of the water is 80°C. How much of the solute will be left undissolved at the bottom of the test tube?
A) 50 grams B) 60 grams C) 70 grams D) 80 grams E) 56 grams

Science Chemistry Unit Solutions Review

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**

Science Chemistry Unit Solutions Review:

Catalogue American International College,1912 **Understanding Chemistry** Herman J. Bruckman,Allan Cruickshanks,1988 Nuclear Science Abstracts ,1974 *Resources in Education* ,1984 **Chemical Principles Student's Study Guide & Solutions Manual** John Krenos,Joseph Potenza,Loretta Jones,Lynn Koplitz,Thomas Spence,2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text **Chemical News and Journal of Physical Science** ,1909 The Chemical News and Journal of Industrial Science ,1913 Catalog of Copyright Entries. Third Series Library of Congress. Copyright Office,1976 *Wastes: Solutions, Treatments and Opportunities III* Candida Vilarinho,Fernando Castro,Margarida Gonçalves,Ana Luísa Fernando,2019-08-08 *Wastes Solutions Treatments and Opportunities III* contains selected papers presented at the 5th edition of the International Conference *Wastes Solutions Treatments and Opportunities* that took place on 3 6 September 2019 in Costa da Caparica Portugal The *Wastes* conference which takes place biennially is a prime forum for sharing innovation technological development and sustainable solutions for the waste management and recycling sectors around the world counting with the participation of experts from academia and industry The papers included in this book cover a wide range of topics including Wastes as construction materials Wastes as fuels Waste treatment technologies MSW management Recycling of wastes and materials recovery Environmental economic and social aspects in waste management Life cycle assessment Circular economy and wastes refineries Logistics policies regulatory constraints and markets in waste management **Chemical News and Journal of Industrial Science** ,1903

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