

Unit 1: Chemistry

Name

TEST

E/Us	/20	T/5	/10	COMM	/15	APP	/5
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Part 1: Multiple Choice

Mark the appropriate letter on the grid below (20 marks- E/Us).

- Which of the following is an example of a physical change?
 - Energy is released or absorbed
 - A precipitate forms
 - A gas is produced
 - Light is emitted
 - Change of state
- Which of the following elements is an example of a halogen?
 - Oxygen
 - Neon
 - Bromine
 - Lithium
 - Magnesium
- Which of the following states of matter indicates that a substance is dissolved in water?
 - ls)
 - l)
 - g)
 - sl)
 - aq)
- In all chemical reactions the:
 - the mass and volume of both the reactants and the products must be equal.
 - the masses of the combining chemicals must be equal.
 - total mass of the reactants equals the total mass of the products
 - the total volume of the reactants equals the total volume of the products
- Al₂O₃ _____ is a pure substance that cannot be broken down further by physical or
 - chemical means
- Which of the following is an example of an ionic compound?
 - NH₃
 - K₂CO₃
 - CaO
 - H₂O
 - CO₂
- When a chemical reaction takes place in a closed system, which of the following statements best describes the principle of conservation of mass?
 - The total mass of the substances before the reaction is greater than the total mass of the substances after the reaction.
 - The total mass of the substances before the reaction is equal to the total mass of the substances after the reaction.
 - The total mass of the substances before the reaction is less than the total mass of the substances after the reaction.
 - The total mass of the reactants decreases, while the total mass of the products increases.
- What type of reaction occurs between calcium and hydrochloric acid (HCl)?
 - synthesis
 - decomposition
 - single displacement
 - double displacement
- _____ is the deterioration of any metal
 - Oxidation

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ML Yell



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