

Chapter 11 Nuclear Chemistry

Review Questions

Chemical energy is energy stored in the bonds of atoms and molecules. Nuclear energy is stored through the nucleus of atoms changing their own different nuclei. Nuclear energy is produced by fission.

The Nucleus



Nucleus is composed of two nucleons: protons and neutrons.
The number of protons is the atomic number.
The number of protons and neutrons together is the mass number.

Isotopes that have the same number of protons but different mass, the number of neutrons is different in their atoms.

Radioactivity is emitted particles or ion elements produced by an unstable nuclei. It can be called radionuclides: ^{22}Na , ^{60}Co , ^{131}I , ^{137}Cs .

Table 11.1: Common isotopes, their half-lives and uses			
Isotope	Half-life	Use	Decay mode
^{14}C	5730 years	Dating	β^-
^{32}P	14.3 days	Tracer	β^-
^{60}Co	5.27 years	Radiotherapy	β^- , γ
^{131}I	8.02 days	Radiotherapy	β^- , γ
^{137}Cs	30.17 years	Radiotherapy	β^- , γ

Equation: Radioactive decay

- Alpha decay is the loss of an α -particle



- Beta decay is the loss of a β -particle or high-speed electrons, produced by the nucleus: $^{14}_6\text{C} \longrightarrow ^{14}_7\text{N}$



- Gamma radiation is the loss of a γ -ray, which is high-energy radiation that released when's transformation after loss of a nuclear particle: $^{60}_{27}\text{Co}$

Review Nuclear Chemistry Key

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