

## Exploring Equilibrium 2.4.10

### How LeQ operates

1. The stability of liquid solutions depends on the position of equilibrium. As defined above, the chemical state of a reaction mixture at the chemical level can be used to describe the position of the system relative to the stability of liquid solutions in the system.
2. Equilibrium constants will be greater relative to the system as the equilibrium constant increases, increasing the temperature profile in the system as well. Thus, the system will favor the position of equilibrium as shift towards the products as temperature rises. The change in the position of equilibrium will be greater as the temperature of the system rises.

### How LeQ operates

1.  $K_{eq}$  is the ratio of the product of the concentrations of the products to the concentration of the reactants. The equilibrium constant is the ratio of the concentration of the products to the concentration of the reactants.
  2.  $K_{eq}$  is the ratio of the product of the concentrations of the products to the concentration of the reactants. The equilibrium constant is the ratio of the concentration of the products to the concentration of the reactants. The change in the position of equilibrium will be greater as the temperature of the system rises. The change in the position of equilibrium will be greater as the temperature of the system rises.
  3.  $K_{eq}$  is the ratio of the product of the concentrations of the products to the concentration of the reactants. The equilibrium constant is the ratio of the concentration of the products to the concentration of the reactants. The change in the position of equilibrium will be greater as the temperature of the system rises. The change in the position of equilibrium will be greater as the temperature of the system rises.
- 100
- (a) When the concentration of the reactants and products are relatively small, the system will favor the position of equilibrium as shift towards the products as temperature rises. The change in the position of equilibrium will be greater as the temperature of the system rises.
1. (a) Adding a species which appears on the right side of an equation will shift the equilibrium to the left side of the equation.
  2. (b) Adding a species which appears on the left side of an equation will shift the equilibrium to the right side of the equation.

# Restoring Balance Lab Answers

**H Kauffman**



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