

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO_3	Chloric acid	H_3O^+	ClO_3^{-1}	Chlorate
H_2SO_3	Sulfurous acid	$2\text{H}_3\text{O}^+$	SO_3^{-2}	Sulfite
H_2SO_4	Sulfuric acid	$2\text{H}_3\text{O}^+$	SO_4^{-2}	Sulfate
H_3PO_3	Phosphorous acid	$3\text{H}_3\text{O}^+$	PO_3^{-3}	Phosphite
H_3PO_4	Phosphoric acid	$3\text{H}_3\text{O}^+$	PO_4^{-3}	Phosphate
HNO_3	Nitric acid	H_3O^+	NO_3^{-1}	Nitrate
HNO_2	Nitrous acid	H_3O^+	NO_2^{-1}	Nitrite
H_2CO_3	Carbonic acid	$2\text{H}_3\text{O}^+$	CO_3^{-2}	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO_3 ?

ClO_3^{-1} has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl , not HClO_3 .

10. Write a rule for naming ternary acids.

- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.

- "ate" is replaced with "ic" and "ite" is replaced with "ous".

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



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Table of Contents Polyatomic Ions Pogil Worksheet Answers

1. Understanding the eBook Polyatomic Ions Pogil Worksheet Answers
 - The Rise of Digital Reading Polyatomic Ions Pogil Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ions Pogil Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polyatomic Ions Pogil Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ions Pogil Worksheet Answers

- Personalized Recommendations
 - Polyatomic Ions Pogil Worksheet Answers User Reviews and Ratings
 - Polyatomic Ions Pogil Worksheet Answers and Bestseller Lists
5. Accessing Polyatomic Ions Pogil Worksheet Answers Free and Paid eBooks
 - Polyatomic Ions Pogil Worksheet Answers Public Domain eBooks
 - Polyatomic Ions Pogil Worksheet Answers eBook Subscription Services
 - Polyatomic Ions Pogil Worksheet Answers Budget-Friendly Options
 6. Navigating Polyatomic Ions Pogil Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Polyatomic Ions Pogil Worksheet Answers Compatibility with Devices
 - Polyatomic Ions Pogil Worksheet Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Polyatomic Ions Pogil Worksheet Answers
 - Highlighting and Note-Taking Polyatomic Ions Pogil Worksheet Answers
 - Interactive Elements Polyatomic Ions Pogil Worksheet Answers
 8. Staying Engaged with Polyatomic Ions Pogil Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polyatomic Ions Pogil Worksheet Answers
 9. Balancing eBooks and Physical Books Polyatomic Ions Pogil Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Polyatomic Ions Pogil Worksheet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Polyatomic Ions Pogil Worksheet Answers
 - Setting Reading Goals Polyatomic Ions Pogil Worksheet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Polyatomic Ions Pogil Worksheet Answers

- Fact-Checking eBook Content of Polyatomic Ions Pogil Worksheet Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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