

Name \_\_\_\_\_

Period \_\_\_\_\_

Date Tuesday, November 1 2016

## Periodic Trends

### ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.  
a) Al B                      b) S O                      c) Br Cl  
d) Na Al                      e) O F                      f) Mg Ca

6. Put the following elements in order from smallest to largest atomic radius **and** explain why:  
C, O, Sn, Sr.

### ELECTRONEGATIVITY

7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.  
a) Ca Ga                      b) Li O                      c) Cl S                      d) Br As                      e) Ba Sr                      f) O S

### GENERAL QUESTIONS

14. Which group tends to form +1 ions? \_\_\_\_\_
15. Which group tends to form +2 ions? \_\_\_\_\_
16. Which group tends to form -1 ions? \_\_\_\_\_
17. Which group tends not to form ions or react? \_\_\_\_\_

# Periodic Trends Worksheet Extension Questions Answers

**Jicheng Xie**



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