

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|-----------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization energy) increases.

Periodic Trends Reactivity Lab Answer Key

Lei Shi

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name 'Lei Shi'.

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book emphasizes the chemical aspects of this science and therefore the mutual reactivity of metals and the characteristics of intermetallic compounds Topics included are Phase diagrams of alloy systems Many intermetallic systems form several compounds generally not obeying common simple stoichiometric rules which are often homogeneous in a certain range of compositions The stability and extension of these phases are conveniently presented through phase diagrams Selected aspects of intermetallics structural chemistry with emphasis on the solid state The general structural characteristics of intermetallic phases are considered with attention to nomenclature and to alternative and complementary methods of presenting crystal chemical data A brief account is given of derivative and degenerate structures modular aspects of crystal structures and of a few special groups of alloys such as quasicrystals and amorphous alloys A number of selected structural prototypes with typical features their possible grouping in structural families and their distribution among different types of alloys are provided Intermetallic reactivity trends in the Periodic Table Attention is given to a few selected elemental parameters such as electron configuration and valence electron number and to their changes along the Table which act as reference factors of the intermetallic behaviour As an example the relationships are considered between crystal structure and the number of valence electrons per atom or per formula in various classes of compounds or solid solution phases Alloying behaviour systematics of intermetallic systems with a description of the intermetallic reactivity of each element or group of elements in the order of their position in the Periodic Table For each pair of metallic elements their capability to form intermediate phases is summarised by maps and schemes A description of small scale preparation methods of intermetallics A number of interesting and significant peculiarities are e g those related to their high melting points insolubility in common solvents etc Systematic treatment of alloying behaviour Wide overview of intermetallic chemistry Illustrated with many examples

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