

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES  
(CHEMISTRY)**

**TABLE 1      PHYSICAL CONSTANTS**

| NAME                            | SYMBOL    | VALUE                                     |
|---------------------------------|-----------|-------------------------------------------|
| Magnitude of charge on electron | $e$       | $1,6 \times 10^{-19} \text{ C}$           |
| Mass of an electron             | $m_e$     | $9,1 \times 10^{-31} \text{ kg}$          |
| Standard pressure               | $p^\circ$ | $1,01 \times 10^5 \text{ Pa}$             |
| Molar gas volume at STP         | $V_m$     | $22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$ |
| Standard temperature            | $T^\circ$ | $273 \text{ K}$                           |
| Avogadro's constant             | $N_A$     | $6,02 \times 10^{23} \text{ mol}^{-1}$    |
| Faraday's constant              | $F$       | $96\,500 \text{ C} \cdot \text{mol}^{-1}$ |

**TABLE 2      CHEMISTRY FORMULAE**

|                                         |                                                                                                                                                                             |                     |
|-----------------------------------------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---------------------|
| $n = \frac{m}{M}$                       | $n = \frac{N}{N_A}$                                                                                                                                                         | $n = \frac{V}{V_m}$ |
| $c = \frac{n}{V}$ OR $c = \frac{m}{MV}$ | $K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$<br>at $25^\circ \text{C}$ (298 K)                                                                    |                     |
| $q = It$<br>$q = nF$                    | $E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$<br>$E_{\text{cell}}^\circ = E_{\text{oxidising agent}}^\circ - E_{\text{reducing agent}}^\circ$ |                     |

# Physical Science P2 Grade1exampler

**Bellaire, Tracy**



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