

Assessment

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Understanding Concepts

11. What is a stable electron configuration?
12. What does each dot in an electron dot diagram represent?
13. What process changes atoms into ions?
14. What keeps the ions in their fixed positions within a crystal lattice?
15. What are subscripts used for in chemical formulas?
16. Explain why a melted ionic compound is a good conductor of electric current, but a solid ionic compound is a poor conductor of electric current.
17. What distinguishes single, double, and triple covalent bonds?
18. Explain why the covalent bonds in molecules of O_2 and N_2 are always nonpolar.
19. Explain why, in a covalent bond between oxygen and hydrogen, the hydrogen atom has a partial positive charge and the oxygen atom has a partial negative charge.
20. What is the name of the binary compound formed from potassium and iodine?
21. Write the formulas for the compounds called copper(I) chloride and copper(II) chloride.
22. Name the compounds represented by the space-filling models labeled A, B, and C.



23. In general, what determines the strength of metallic bonds?
24. What properties of copper and tin change when these metals are mixed together to form bronze?
25. What advantage of magnesium is retained in magnesium alloys? What disadvantage is reduced?

Chemical Bonds 187

Assessment



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Reviewing Content

- | | | |
|-------|------|------|
| 1. b | 2. c | 3. a |
| 4. d | 5. c | 6. b |
| 7. d | 8. b | 9. c |
| 10. a | | |

Understanding Concepts

11. In a stable electron configuration, the highest occupied energy level is filled with electrons.
12. Each dot represents a valence electron.
13. The transfer of electrons.
14. Attractions between neighboring cations and anions keep the ions in fixed positions within the lattice.
15. A subscript is used to show the number of atoms of an element in a molecule or the ratio of ions in a crystal lattice.
16. When an ionic compound melts, ions can move away from their fixed locations in the crystal lattice.
17. Two atoms share two electrons in a single bond, four in a double bond, and six in a triple bond.
18. The covalent bonds in molecules of elements are always nonpolar because the atoms have the same attraction for electrons.
19. The oxygen atom has a greater attraction for electrons than the hydrogen atom does.
20. Potassium iodide
21. CuCl and CuCl_2
22. A is sulfur trioxide, B is carbon monoxide, and C is nitrogen dioxide
23. In general, the more valence electrons a metal has, the stronger the metal bonds are.
24. A mixture of copper and tin is harder and stronger than either metal in its pure form.
25. The advantage that is retained is that magnesium is a lightweight metal. The disadvantage that is reduced is that magnesium is a soft metal.

Index

- atoms
11–16, 26, 30, 33
8, 17–19, 27, 34–35, 39–40, 42
20–22, 28–29, 31–32, 36–38
23–25, 41

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Table of Contents Physical Science Chapter 4 Review Answers Pearson

1. Understanding the eBook Physical Science Chapter 4 Review Answers Pearson
 - The Rise of Digital Reading Physical Science Chapter 4 Review Answers Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 4 Review Answers Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 4 Review Answers Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 4 Review Answers Pearson
 - Personalized Recommendations
 - Physical Science Chapter 4 Review Answers Pearson User Reviews and Ratings

- Physical Science Chapter 4 Review Answers Pearson and Bestseller Lists
- 5. Accessing Physical Science Chapter 4 Review Answers Pearson Free and Paid eBooks
 - Physical Science Chapter 4 Review Answers Pearson Public Domain eBooks
 - Physical Science Chapter 4 Review Answers Pearson eBook Subscription Services
 - Physical Science Chapter 4 Review Answers Pearson Budget-Friendly Options
- 6. Navigating Physical Science Chapter 4 Review Answers Pearson eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 4 Review Answers Pearson Compatibility with Devices
 - Physical Science Chapter 4 Review Answers Pearson Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 4 Review Answers Pearson
 - Highlighting and Note-Taking Physical Science Chapter 4 Review Answers Pearson
 - Interactive Elements Physical Science Chapter 4 Review Answers Pearson
- 8. Staying Engaged with Physical Science Chapter 4 Review Answers Pearson
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 4 Review Answers Pearson
- 9. Balancing eBooks and Physical Books Physical Science Chapter 4 Review Answers Pearson
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 4 Review Answers Pearson
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chapter 4 Review Answers Pearson
 - Setting Reading Goals Physical Science Chapter 4 Review Answers Pearson
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chapter 4 Review Answers Pearson
 - Fact-Checking eBook Content of Physical Science Chapter 4 Review Answers Pearson
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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