

Periodicity Chemistry Worksheet

A. Periodic table

1. Which are metals? Circle your answers: C, Na, F, Cs, Ba, Ni

Which metal in the list above has the most metallic character? Explain.

Cs - It is the metal with the ↑ shielding & ↓ ENC

2. Write the charge that each of the following atoms will have when it has a complete set of valence electrons forming an ion.

O⁻²

Na⁺¹

F⁻¹

N⁻³

Ca⁺²

Ar⁰

3. What is the most common oxidation number for calcium? Explain.

Ca⁺² - Calcium has 2 valence electrons, it will lose 2 electrons

Name two more elements with that oxidation number and explain your choice.

Any of the two: Be, Mg, Se, Ba, Ra - Same valence electrons / same family

4. What element in period 3 is a metalloid? Silicon

5. When element with atomic number 118 is discovered, what family will it be in? Noble Gas

Use the following word bank for questions 6 through 13

(Obviously, you will need to use several words more than once! ☹)

Alkali metals

Alkaline earth metals

Halogens

Noble gases

Transition metals

Noble gases

6. The Alkali metals have a single electron in the highest energy level.
7. The Alkaline earth metals achieve the electron configurations of noble gases by losing two electrons.
8. The transition metals are metals that can hold up to 10 electrons in their sublevel shape
9. The halogens achieve the electron configuration of noble gases by gaining one electron.
10. The noble gases have full s and p orbitals in the highest occupied energy levels.
11. The noble gases are stable and un-reactive.
12. The halogens are highly reactive and readily form salts with metals.
13. The Alkaline earth metals are metals that are more reactive than the transition elements but less reactive than the alkali metals.
14. Predict the oxidation number based on the electron configuration shown.
- | | | | |
|-----------------------|------------------------|-----------------------|------------------------|
| $1s^2 2s^2 2p^6 3s^2$ | <u>Mg⁺²</u> | $1s^2 2s^2 2p^6 3s^1$ | <u>Na⁺¹</u> |
| $1s^2 2s^2 2p^6$ | <u>Ne⁰</u> | $1s^2 2s^2 2p^5$ | <u>F⁻¹</u> |

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M Woodhall



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