

# User Manual

## S860 | S760

Thank you for buying a Samsung Camera.  
This manual will guide you through using the camera, including capturing images, downloading images and using the application software.  
Please read this manual carefully before using your new camera.

※ This manual is based on model S860.

ENGLISH

### Instructions

Use this camera in the following order.



### Getting to Know your camera

Thank you for buying a Samsung Digital Camera.

- Prior to using this camera, please read the user manual thoroughly.
- If you require After Sales service, please bring the camera and the cause of the camera malfunction (such as battery, Memory card etc.) to the A/S centre.
- Please check that the camera is operating properly prior to when you intend to use it (e.g., for a trip or an important event) to avoid disappointment. Samsung takes no responsibility for any loss or damage that may result from camera malfunction.
- Keep this manual in a safe place.
- If you use a card reader to copy the images on the memory card to your PC, the images could be damaged. When transferring the images taken with the camera to your PC, be sure to use the supplied USB cable to connect the camera to your PC. Please note that the manufacturer holds no responsibility for loss or damage of images on the memory card due to the use of a card reader.
- The contents and illustrations of this manual are subject to change without prior notice for upgrading the camera functions.

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# Samsung S860 Owners Manual

**Siddappa N.Byrareddy**



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web aug 15 2020 the equilibrium formulated in scheme 1 between hydrogen ion and the corresponding acid and base can be called a simple acid base equilibrium by mixing two simple systems a double acid base system and an acid base equilibrium result that can always be formulated as follows  $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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web the  $\text{A}^-$  ion is a base because it is accepting a proton from the  $\text{H}_3\text{O}^+$  the reversible reaction contains two acids and two bases we think of them in pairs called conjugate pairs when the acid  $\text{HA}$  loses a proton it forms a base  $\text{A}^-$  when the base  $\text{A}^-$  accepts a proton back again it obviously reforms the acid  $\text{HA}$

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