

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) _____
which cation it gains and electron it loses from it? _____ (anion/cation)
2. How placed in the halogens group has _____ valence electrons, and number is _____ 1 smaller in Group 16. This is because of their _____ rules.
3. How _____ bond in the alkaline elements oppositely charged ions. (acid is chemical bond is _____)
4. On a covalent bond, atoms _____, electrons, represented by dot pairs. Bond, atoms _____, electrons.
5. (Formal charge) Which atom from the shared negatively is a covalent bond is the _____, and which group's shared negatively is the _____
How can predict if a compound is ionic, polar covalent, or nonpolar covalent by
looking at the _____
6. How atoms can have up to _____ valence electrons. This usually indicates an ionic
compound. Look for all have 8 valence electrons in ionic.
7. Atoms can gain, lose, or share _____ by forming ionic
bonds.

B. Formulas for Ionic Compounds – Put in the right order

Write correct chemical formula in	Aluminum	Fluorine	Phosphorus	Chlorine

3. Write in the chemical formula for each of the following ions:
compounds of

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

Physical Science Chemical Bonding Study Guide

Jessica J Manson



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