Chemistry I: Periodicity Worksheet

1) Arrange the following elements in order of increasing size: Al, B, C, K, and Na.

C B Al Na K smallest largest

Arrange the following elements in order of increasing size: Ca, Rb, P, Ge, and Sr.

P Ge Ca Sr Rb smallest largest

Circle the atom or ion in each pair that has the larger radius.

- a) Clor CI
- b) All or O
- c) In or I

4) Select the atom or ion in each pair that has the larger radius.

- a) Cs or Rb
- b) O³ or O
- c) BrorAs

5) Which of the following groups of elements is arranged correctly in order of increasing ionization energy? (Answer by circling the correct answer.)

Non

- a) $C \le Si \le Li \le Ne$
- b) Ne \leq Si \leq C \leq Li
- e) Li < Si < C < Ne
- d) Ne < C < Si < Li</p>

6) Arrange the following atoms in order of increasing ionization energy: Li, K, C, and N.

K LI C N lowest highest

Arrange the following atoms in order of increasing ionization energy: Si, K, As, and Ca.

K Ca As Si lowest highest

8) Compare the elements Na, Mg, O, and P.

- a) Which has the largest atomic radius?
- b) Which has the highest electron affinity? O

c) Place the elements in order of increasing ionization energy.

Na Mg P O lowest highest

Periodicity Worksheet Answer Key

Pam Walker, Elaine Wood

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