

## 5 - Periodicity Activity

2023/24

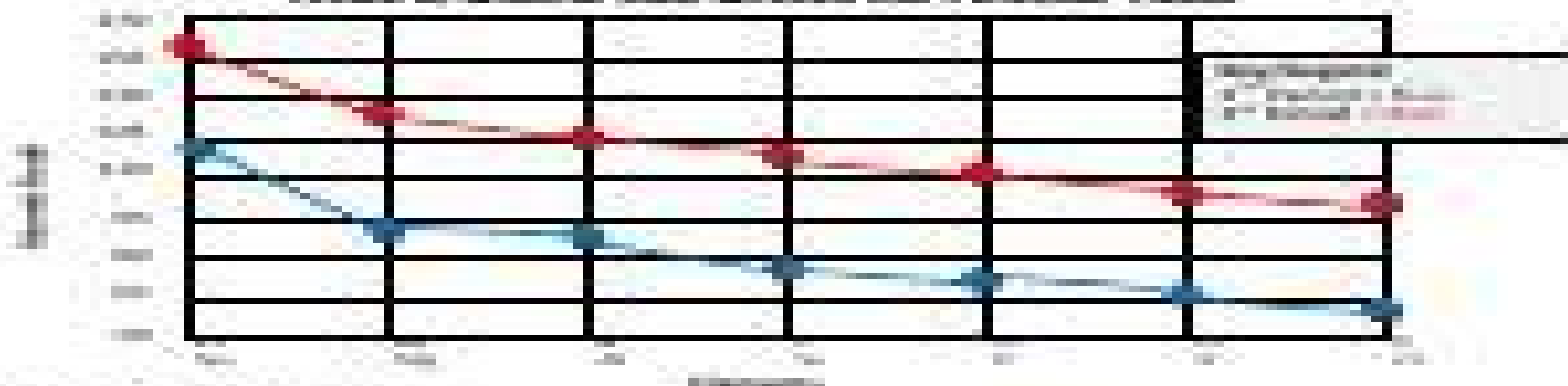
Instructions: Using the graphs above fill the boxes of atomic properties below.

1. Choose which graph below best fits graphing Period 2's data points (graph below)
2. Is a slight downward curve or does graph Period 2's data fit the other graph?
3. Label the line as ionisation enthalpy
4. downward slope/curve.

	1s	2s	2p	3s	3p	3d	4s	4p
1	1s	2s	2p	3s	3p	3d	4s	4p
2	2s	2p	3s	3p	3d	4s	4p	4d
3	3s	3p	3d	4s	4p	4d	4f	5s
4	4s	4p	4d	4f	5s	5p	5d	5f
5	5s	5p	5d	5f	6s	6p	6d	6f
6	6s	6p	6d	6f	7s	7p	7d	7f

Now label the elements from H<sub>1</sub> to H<sub>10</sub> on the

Trend in Atomic Size across the Periodic Table



Illustrated Trend Equations

1. Following the trend in atomic size across the periodic table decreases
2. This causes two more than just more electrons protons/electrons they give all electrons pulling themselves closer to the nucleus together

Worked Trend Equations

1. Periodic trend of electronegativity is increases (e.g. Na, Cl, Br, F, O). The less electronegative elements on your graph are on the left of the periodic table decreases
2. This causes two more than just more electrons protons/electrons (e.g. give all electrons) pulling them closer together

# Periodicity Activity Okstate Answer Key

**Raffaela Di Napoli**



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