

## Study Guide: Periodic Table and Periodic Trends (Honors)

1. Differentiate between Mendeleev's periodic table and Moseley's periodic table.

**Mendeleev put the elements in order by atomic mass, but he noticed that in several places, the atomic masses did not always increase sequentially. Moseley discovered that ordering by atomic number was the correct way to describe the order.**

2. State the periodic law.

**The properties of the elements are periodic functions of their atomic numbers.**

3. Vertical columns on the periodic table are called **groups or families** and horizontal rows are called **periods**.

4. Elements with similar chemical properties are placed in the same **group**.

5. Complete the table below.

| Electron Config.   | Block | Period | Group/Type           | Element   |
|--|-------|--------|----------------------|-----------|
| [Xe] 6s <sup>2</sup> 4f <sup>14</sup> 5d <sup>10</sup> 6p <sup>6</sup> | p     | 6      | 18 / Noble Gas       | Radon     |
| [Xe] 6s <sup>1</sup>   | s     | 6      | 1 / Alkali Metal     | Cesium    |
| [Ar] 4s <sup>1</sup> 3d <sup>5</sup>                                   | d     | 4      | 6 / Trans. Metal     | Chromium  |
| [He] 2s <sup>2</sup>   | s     | 2      | 2 / Alk. Earth Metal | Beryllium |
| [Kr] 5s <sup>1</sup> 4d <sup>10</sup>                                  | d     | 5      | 11 / Trans. Metal    | Silver    |
| [He] 2s <sup>2</sup> 2p <sup>5</sup>                                   | p     | 2      | 17 / Halogen         | Fluorine  |

6. Determine whether the following statements are describing alkali metals (A), alkaline earth metals (AE), transition metals (TM), halogens (H), noble gases (NG):

- a) most reactive non-metals **H**
- b) s<sup>2</sup>p<sup>6</sup> valence configuration **NG**
- c) second most reactive metals **AE**
- d) s<sup>1</sup> valence configuration **A**
- e) exceptions found to e- configurations **TM**
- f) soft metals **A**
- g) d orbitals being filled **TM**

7. Would you expect Li and S to have different chemical properties? Why? **Yes; Li is an alkali metal with 1 valence e-, and S is a nonmetal with 6 valence e-.**

8. Explain why a magnesium atom is smaller than: a sodium atom; a calcium atom. **Mg is smaller than Na because the greater nuclear charge in Mg pulls the e- cloud in tighter (periodic trend). Mg is smaller than Ca because it has fewer energy levels (group trend).**

9. Predict the size of the astatine atom compared to the sulfur atom. **At is larger than S, because At is 3 periods down from S, but only 1 group over. The trends compete, but the sizeable increase in energy levels overcomes the slight difference in nuclear charge.**

10. Would you expect a Cl<sup>-</sup> ion to be larger or smaller than an Mg<sup>2+</sup> ion. Explain. **A Cl<sup>-</sup> ion is larger than a Mg<sup>2+</sup> ion, because the chloride ion has 3 energy levels, matching**

# Periodic Trends Study Guide

**Merrill**



## Periodic Trends Study Guide:

**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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